

# Step 8. Periodic table trends

Electronegativity

Ionization energy

He

Atomic radius

The periodic table displays various elements with their symbols and atomic numbers. Overlaid on the table are several arrows indicating trends:

- Electronegativity:** A black arrow pointing from left to right across the top of the table.
- Ionization energy:** A green arrow pointing from left to right across the top of the table, and a green arrow pointing from bottom to top on the right side of the table.
- Atomic radius:** A red arrow pointing from top to bottom on the left side of the table, and a red arrow pointing from right to left across the bottom of the table.
- Non-metallic properties:** A blue arrow pointing from bottom-left to top-right, passing through the upper right quadrant of the table.
- Metallic properties:** A light blue arrow pointing from top-left to bottom-right, passing through the lower left quadrant of the table.

Ionization Energy

Electronegativity

Atomic radius

An arrow head shows the INCREASE of the property

Example: comparing K atom with Br atom

- Atomic radius of K is HIGHER than atomic radius of Br
- Ionization energy of K is LOWER than ionisation energy of Br
- Metallic properties of K are HIGHER than metallic properties of Br



# Periodic Trends Key

**JE Gale**



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## **Table of Contents Periodic Trends Key**

1. Understanding the eBook Periodic Trends Key
  - The Rise of Digital Reading Periodic Trends Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Key
  - Personalized Recommendations
  - Periodic Trends Key User Reviews and Ratings
  - Periodic Trends Key and Bestseller Lists
5. Accessing Periodic Trends Key Free and Paid eBooks
  - Periodic Trends Key Public Domain eBooks
  - Periodic Trends Key eBook Subscription Services
  - Periodic Trends Key Budget-Friendly Options

6. Navigating Periodic Trends Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Key Compatibility with Devices
  - Periodic Trends Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Key
  - Highlighting and Note-Taking Periodic Trends Key
  - Interactive Elements Periodic Trends Key
8. Staying Engaged with Periodic Trends Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Key
9. Balancing eBooks and Physical Books Periodic Trends Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Key
  - Setting Reading Goals Periodic Trends Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Key
  - Fact-Checking eBook Content of Periodic Trends Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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