

# 6

## THE PERIODIC TABLE

### Chapter 6 Review Package - Part 1

#### A. Matching

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A	Column B
_____ 1. metal	a. positively charged ion
_____ 2. halogen	b. an element whose highest occupied <i>s</i> or <i>p</i> sublevels are partially filled
_____ 3. inner transition metal	c. <del>the energy needed to remove an electron from an atom in the gaseous state</del>
_____ 4. representative element	d. good conductor of heat and electric current
_____ 5. <del>ionization energy</del>	e. an element whose highest occupied <i>s</i> sublevel and a nearby <i>d</i> sublevel contain electrons
_____ 6. alkaline earth metal	f. any nonmetal in Group 7A
_____ 7. cation	g. <del>the tendency of an atom to attract electrons when the atom is in a compound</del>
_____ 8. electronegativity	h. an element in which the highest occupied <i>s</i> and <i>p</i> sublevels are filled
_____ 9. transition metal	i. an element in Group 2A
_____ 10. noble gas	j. an element whose highest occupied <i>s</i> sublevel and a nearby <i>f</i> sublevel contain electrons

#### B. Multiple Choice

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. In the periodic table, there is a periodic pattern in the physical and chemical properties of elements when they are arranged in order of
- increasing atomic mass.
  - increasing electronegativity.
  - increasing atomic radius.
  - increasing atomic number.



# Periodic Trends Ch 6 Answers

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