

NAME \_\_\_\_\_

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### GRAPHING PERIODIC TRENDS

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.

Define the following terms(starting on page 170):

Atomic radius-

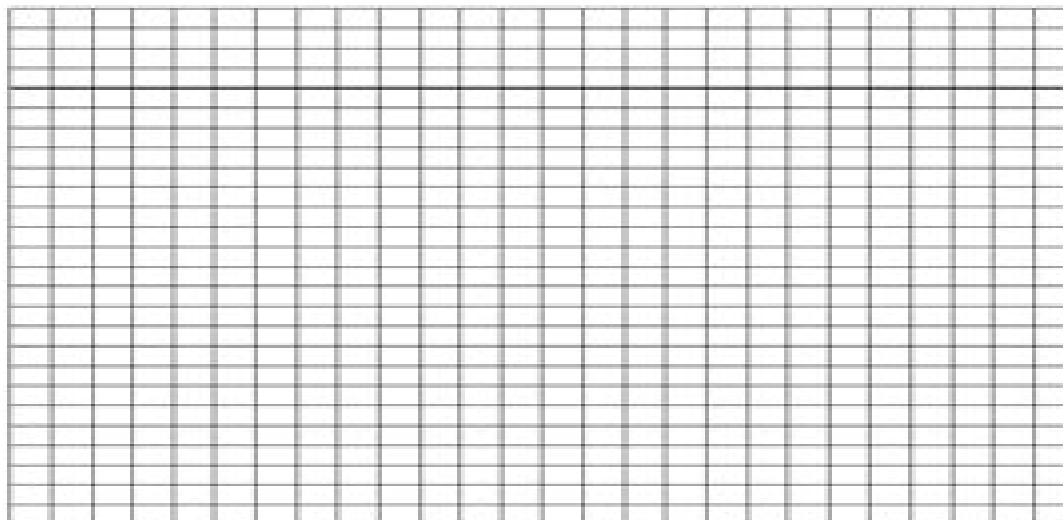
First ionization energy-

Electronegativity-

#### Graph 1 – Atomic Radius as a function of Atomic Number

- A. Create a graph of the atomic radius as a function of atomic number. Plot atomic number on the X axis and atomic radius on the Y axis. Remember to label the axes!
- B. Use a colored pen, pencil or highlighter to **trace** over the element's period (horizontal row on the periodic table). For example: use GREEN to **trace** for all of the elements in row 1, then use YELLOW to **trace** for all of the elements in row 2, then use ORANGE to **trace** for all the elements in row 3, then use BLUE to **trace** for all the elements in row 4.

Symbol	Atomic Number	Atomic Radius(pm)
H	1	37
He	2	50
Li	3	123
Be	4	89
B	5	80
C	6	77
N	7	70
O	8	66
F	9	64
Ne	10	67
Na	11	157
Mg	12	136
Al	13	125
Si	14	117
P	15	110
S	16	104
Cl	17	99
Ar	18	98
K	19	203
Ca	20	174



1. Describe the trend in atomic radius as the atomic number increases across a period.
2. Describe the trend in atomic radius as the atomic number increases down a group.

# Periodic Trends Answers Extension Questions

**Tina Overton, Jonathan Rourke, Fraser  
A. Armstrong**



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## **Table of Contents Periodic Trends Answers Extension Questions**

1. Understanding the eBook Periodic Trends Answers Extension Questions
  - The Rise of Digital Reading Periodic Trends Answers Extension Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Answers Extension Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Answers Extension Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Answers Extension Questions
  - Personalized Recommendations
  - Periodic Trends Answers Extension Questions User Reviews and Ratings

- Periodic Trends Answers Extension Questions and Bestseller Lists
- 5. Accessing Periodic Trends Answers Extension Questions Free and Paid eBooks
  - Periodic Trends Answers Extension Questions Public Domain eBooks
  - Periodic Trends Answers Extension Questions eBook Subscription Services
  - Periodic Trends Answers Extension Questions Budget-Friendly Options
- 6. Navigating Periodic Trends Answers Extension Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Answers Extension Questions Compatibility with Devices
  - Periodic Trends Answers Extension Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Answers Extension Questions
  - Highlighting and Note-Taking Periodic Trends Answers Extension Questions
  - Interactive Elements Periodic Trends Answers Extension Questions
- 8. Staying Engaged with Periodic Trends Answers Extension Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Answers Extension Questions
- 9. Balancing eBooks and Physical Books Periodic Trends Answers Extension Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Answers Extension Questions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Answers Extension Questions
  - Setting Reading Goals Periodic Trends Answers Extension Questions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Answers Extension Questions
  - Fact-Checking eBook Content of Periodic Trends Answers Extension Questions
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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