

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

5. a. O C Be Ne Same Period

- b. Na Rb Fr H Same Group

- c. Pb C Sn Si Same Group

- d. Au W S Fr Ne Zn Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- a. O C Be Ne Same Period

- b. Na Rb Fr H Same Group

- c. Pb C Sn Si Same Group

- d. Au W S Fr Ne Zn Challenge

Periodic Trends Worksheet Answers

Mary Louise Turgeon



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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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Table of Contents Periodic Trends Worksheet Answers

1. Understanding the eBook Periodic Trends Worksheet Answers
 - The Rise of Digital Reading Periodic Trends Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Worksheet Answers
 - Personalized Recommendations
 - Periodic Trends Worksheet Answers User Reviews and Ratings
 - Periodic Trends Worksheet Answers and Bestseller Lists
5. Accessing Periodic Trends Worksheet Answers Free and Paid eBooks

- Periodic Trends Worksheet Answers Public Domain eBooks
 - Periodic Trends Worksheet Answers eBook Subscription Services
 - Periodic Trends Worksheet Answers Budget-Friendly Options
6. Navigating Periodic Trends Worksheet Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Trends Worksheet Answers Compatibility with Devices
 - Periodic Trends Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Trends Worksheet Answers
 - Highlighting and Note-Taking Periodic Trends Worksheet Answers
 - Interactive Elements Periodic Trends Worksheet Answers
8. Staying Engaged with Periodic Trends Worksheet Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Worksheet Answers
9. Balancing eBooks and Physical Books Periodic Trends Worksheet Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Worksheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Worksheet Answers
- Setting Reading Goals Periodic Trends Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Worksheet Answers
- Fact-Checking eBook Content of Periodic Trends Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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