

Quiz & Worksheet - Determining the Chemical Formula of Hydrates from Empirical Data

1. What is the difference between a hydrate and an anhydrate?

- ☐ An anhydrate is a hydrate with the water removed.
- ☐ A hydrate is an anhydrate with the water removed.
- ☐ A hydrate is necessary for producing water molecules, but an anhydrate isn't.
- ☐ An anhydrate is necessary for producing water molecules, but a hydrate isn't.

2. In the formula $\text{MgSO}_4 \cdot n\text{H}_2\text{O}$, what does the n before the H_2O stand for?

- ☐ The number of moles of water.
- ☐ The number of moles of hydrogen.
- ☐ The number of moles of hydrate.
- ☐ The number of moles of anhydrate.

3. If there are three moles of H_2O for every one mole of LiClO_4 , then what is the hydrate formula?

- ☐ $\text{LiClO}_4 \cdot 3\text{H}_2\text{O}$
- ☐ $3\text{LiClO}_4 \cdot 3\text{H}_2\text{O}$
- ☐ $\text{LiClO}_4 \cdot (3\text{H}_2\text{O})_3$
- ☐ $3(\text{LiClO}_4) \cdot 3\text{H}_2\text{O}$

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