

## Specific Heat

**DIRECTIONS:** Use formula  $q = (m)(\Delta T)(C)$  to solve the following problems. Show all work and units. Use <http://weebly.com> for specific heat of various substances.

1. A 30g piece of metal absorbs 1,200 joules of heat energy and its temperature change from 25°C to 175°C. Calculate the specific capacity of the metal. What is the likely metal?
2. How many joules of heat are needed to raise the temperature of 15.0 g of iron from 11°C to 68°C, if the specific heat of iron is 0.45 J/g°C?
3. To what temperature will a 50.0 g piece of brass raise if it absorbs 6 kilojoules of heat and its specific heat capacity is 0.38 J/g°C? The initial temperature of the brass is 20.0°C.
4. Identify the substance if 1500.0 g of the substance absorbs 74,500 joules of heat, and its temperature changes from 25°C to 47°C.
5. 100.0 g of 1.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.
6. 25.0 g of antimony is heated from 20°C to 160°C, and absorbs 750 joules of heat in the process. Calculate the specific heat capacity of antimony.
7. What is the specific heat capacity of quartz glass if 200.0 g of the glass absorbs 2520 joules of heat and the temperature rises 18.0°C?
8. If a sample of titanium is initially at 25°C, what is its final temperature if 250.0 g of titanium absorbs 1.0 kilojoules of heat, and the specific heat of titanium is 0.523 J/g°C?
9. How much energy must be absorbed by 300.0 g of steel to increase its temperature from 100.0°C to 450.0°C? (C of steel = .48 J/g °C)
10. When 55.0 g of platinum drops its temperature from 300.0°C to 250.0°C, how much heat energy is released? (C of platinum = .133 J/g °C)
11. How much energy is required to heat 125.0 g of lithium from 2.0°C to 30.0°C? (C of lithium = 3.6 J/g°C)
12. How much heat (in kJ) is given out when 350.0 g of lead cools from 200.0°C to 30.0°C?

# Practice Problems For Specific Heat

**Morris Hein, Susan Arena**



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