

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, S, Cl, O

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, Cl, O

3. Which atom has a higher electronegativity: fluorine or sulfur?

Fluorine is more electronegative than sulfur because it is further to the right.

4. Which two elements in the same family generally have similar properties?

There have been two atoms that share a same property such as potassium.

5. Arrange **fluorine** the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (decreasing to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms of the greatest electronegativity? Which element has the lowest?

Fluorine, period 2, electronegativity increases as you go to the rightmost element.

7. Which trend in ionization energy values across a period and the periodic table? Which element has the highest?

Electronegativity increases as you go to the rightmost element.

8. Write the atoms in each pair that has the largest atomic radius.

a. **Br** or **Cl**
 b. **Na** or **K**
 c. **Li** or **Cl**
 d. **Li** or **Br**
 e. **Li** or **K**
 f. **Br** or **K**

a. **Br** or **Cl**
 b. **Na** or **K**
 c. **Li** or **Cl**
 d. **Li** or **Br**
 e. **Li** or **K**
 f. **Br** or **K**

9. Circle the atom in each pair that has the greatest electronegativity.

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