PHYSICAL CHEMISTRY

(Complete Syllabus)

SECTION - A

Straight Objective Type

This section contains 220 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

Choose the correct answer :

1. The mass of oxygen required to produce enough CO which completely reduces x gm of Fe_2O_3. Suppose molecular weight of O_2 = y, molecular weight of Fe_2O_3 = z

   (1) $\frac{2xz}{3y}$ (2) $\frac{3xy}{2z}$

   (3) $\frac{2y}{3xz}$ (4) $\frac{3z}{2xy}$

2. The number of lone pair of electrons present in 2.56 gm of orthorhombic sulphur will be

   (1) 0.02 N_0 (2) 0.16 N_0

   (3) 0.32 N_0 (4) 0.08 N_0

3. A 10.0 g of sample of mixture of calcium chloride and sodium chloride is treated with Na_2CO_3 to precipitate the calcium as calcium carbonate. This $CaCO_3$ is heated to convert all calcium to CaO and final mass of CaO is 1.62 gms. The % by mass of $CaCl_2$ in the original mixture is

   (1) 15.2%
   (2) 32.1%
   (3) 21.8%
   (4) 11.07%

4. 1000 ml of oxygen at STP were passed through an ozoniser and the resulting volume was 888 ml at STP. This quantity of ozonised oxygen is passed through excess of KI solution. The weight of I_2 liberated is

   (1) 25.4 g (2) 12.7 g

   (3) 2.54 g (4) 1.27 g

5. x_1 ml solution having normality 1 is diluted to get a solution of normality y_2. What would be volume required for dilution?

   (1) $x_1\left(1-\frac{y_1}{y_2}\right)$ (2) $x_1\left(\frac{y_2}{y_1}-1\right)$

   (3) $x_1\left(\frac{y_1}{y_2}-1\right)$ (4) $x_1\left(1+\frac{y_1}{y_2}\right)$

6. 1.625 g of an interhalogen compound ICl_x was subjected to electric discharge which split it into constituents and brought to STP. It occupied 112 ml. The interhalogen compound is

   (1) ICl (2) ICl_3

   (3) ICl_5 (4) ICl_7

7. If 12.25 g of a sample of $KClO_3$ of 75% purity and rest vapourisible impurties is strongly heated to give KCl, $KClO_4$ and 0.32 gm O_2. The mass of residue obtained is

   (1) 1.769 g (2) 7.096 g

   (3) 5.327 g (4) 8.85 g

8. A 2 g sample of xenon reacts with fluorine. The mass of compound produced is 3.158 g. The empirical formula of the compound is

   (1) XeF_2 (2) XeF_4

   (3) XeF_5 (4) XeF_6

9. 1 g of carbonate of metal was dissolved in 25 ml of 1 N HCl. The resulting liquid required 5 ml of 1 N NaOH for neutralisation. The equivalent weight

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