

**Concept-Development
Practice Page**

32-1

Coulomb's Law

1. The diagram is of a hydrogen atom.

- Label the proton in the nucleus with a $+$ sign and the orbital electron with a $-$ sign.
- The electrical interaction between the nucleus and the orbital electron is a force of _____ (attractive) _____ (repulsive).
- According to Coulomb's law,

$$F = k \frac{q_1 q_2}{r^2}$$

if the charge of either the nucleus or the orbital electron were greater, the force between the nucleus and the electron would be _____ (greater) _____ (less).

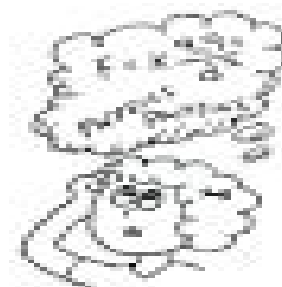
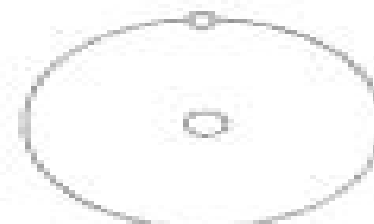
_____ (greater) _____ (less).

and if the distance between the nucleus and electron were greater the force would be _____ (greater) _____ (less).

If the distance between the nucleus and electron were doubled, the force would be _____ (1/4 as much) _____ (1/2 as much) _____ (two times as much) _____ (four times as much).

2. Consider the electric force between a pair of charged particles a certain distance apart. By Coulomb's law:

- If the charge on one of the particles is doubled, the force is _____ (unchanged) _____ (halved) _____ (doubled) _____ (quadrupled).
- If instead the charge on both particles is doubled, the force is _____ (unchanged) _____ (halved) _____ (doubled) _____ (quadrupled).
- If instead the distance between the particles is halved, the force is _____ (unchanged) _____ (halved) _____ (doubled) _____ (quadrupled).
- If the distance is halved and the charge of both particles is doubled, the force is _____ times as great.



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Wolfgang Guggemos

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