Stoichiometry Worksheet and Key

2 KClO₃ → 2 KCl + 3 O₂

How many moles of O₂ will be formed from 1.65 moles of KClO₃?

 $\begin{array}{c|cccc}
1.65 & \text{mol KClO}_3 & \underline{\qquad} & \text{mol O}_2 \\
& \underline{\qquad} & \text{mol KClO}_3 & \underline{\qquad} & \text{mol O}_2
\end{array}$

2. How many moles of KClO₃ are needed to make 3.50 moles of KCl?

3.50mol KCl = ___ mol KClO₃

3. How many moles of KCl will be formed from 2.73 moles of KClO₃?

4 Fe + 3 O₂ → 2 Fe₂O₃

4. How many moles of Fe₂O₃ are produced when 0.275 moles of Fe is reacted?

0.275 mol Fe = ___ mol Fe₂O₃

5. How many moles of Fe₂O₃ are produced when 31.0 moles of O₂ is reacted?

6. How many moles of O₂ are needed to react with 8.9 moles of Fe?

 $2 H_2O \rightarrow 2 H_2 + O_2$

7. How many moles of O₂ are produced when 1.26 moles of H₂O is reacted?

How many moles of H₂O are needed to produce 55.7 moles of H₂?

If enough H₂O is reacted to produce 3.40 moles of H₂, then how may moles of O₂ must have been made?
 (a bit challenging, but just think about it and you can probably figure it out)

Practice Test Stoichiometry Answers

Huangqi Zhang

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