



**NATIONAL
SENIOR CERTIFICATE**

GRADE 11

NOVEMBER 2020

**PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR**

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use permanent marker for writing and drawing.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and ΔH_{fusion} for water in kJ mol⁻¹.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	162
B	102	17

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 In what phase was compound A and B at 0 °C and 100 °C and 101 °C of boiling compound? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What kind of the compound A and B is 1-hexanol? Explain how you proposed these? (2)
- 3.1.5 Which compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Give a reason for your answer. (2)
- 3.1.7 Which compound (A or B) will have the higher vapour pressure? (2)
- 3.1.8 Give a reason for the answer to 3.1.7 by referring to the space for table. (2)

Compound B (H₂O) will show in the table has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 100 °C.

- 3.1.8 Explain why the boiling point of compound B is 100 °C? (2)
- 3.1.9 Explain the difference in the boiling point of compound A and B by referring to the ΔH_{vap} and intermolecular or intramolecular forces involved. (2)

3.2 Substances C1 is added to water whereas D is added to water. Substances C1 and D are shown in separate test tubes as shown in the diagram below.



- 3.2.1 In which test tube (C1 or D), will the solid dissolve? (2)
- 3.2.2 Explain the answer to QUESTION 3.2.1 about the solubility in the intermolecular forces involved. (2)

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L Towne

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape with a gradient, partially overlapping the bar's end.

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sadly overlooked In so doing the book will shed much new light on the scientific culture of the late nineteenth and early twentieth centuries and the development of the field of chemistry in Britain

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