

How to Pass the Chemistry Regents

85	Acid/Base	Neutralization	Acid + Base → Water + Salt (4C, Table A) + (20C, Table L) + (24C) + (26C) + (27C) + (28C)	1.24
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25 Which type of reaction occurs when an Arrhenius acid reacts with an Arrhenius base to form a salt and water?

- (1) combustion (2) neutralization
(3) decomposition (4) saponification

26 Which type of reaction occurs when $\text{H}^+(\text{aq})$ reacts with $\text{OH}^-(\text{aq})$?

- (1) combustion (2) fermentation
(3) decomposition (4) neutralization

27 Which substance always forms when an Arrhenius acid reacts with an Arrhenius base?

- (1) CO_2 (3) CH_3OH
(2) H_2 (4) H_2O

28 In a neutralization reaction, an aqueous solution of an Arrhenius acid reacts with an aqueous solution of an Arrhenius base to produce

- (1) an ether and water
(2) an ether and an alcohol
(3) a salt and water
(4) a salt and an alcohol

45 Which equation represents a neutralization reaction?

- (1) $4\text{HClO}_2 \rightarrow 4\text{HCl} + 2\text{HClO}_4$
(2) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
(3) $\text{Ca(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{CaSO}_4 + 2\text{H}_2\text{O}$
(4) $\text{Ba(OH)}_2 + \text{Ca(NO}_3)_2 \rightarrow \text{Ba(NO}_3)_2 + \text{Ca(OH)}_2$

26 Which equation represents neutralization?

- (1) $4\text{Li(s)} + \text{N}_2(\text{g}) \rightarrow 2\text{Li}_3\text{N(s)}$
(2) $2\text{Mg(s)} + \text{O}_2(\text{g}) \rightarrow 2\text{MgO(s)}$
(3) $2\text{KOH(aq)} + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{K}_2\text{SO}_4(\text{aq}) + 2\text{H}_2\text{O(l)}$
(4) $\text{Pb(NO}_3)_2(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow 2\text{KNO}_3(\text{aq}) + \text{PbCrO}_4(\text{s})$

64 Complete the equation in your answer booklet for the neutralization reaction by writing a formula for each product. [1]



83 Complete the equation in your answer booklet by writing the chemical formula for each product. [1]



79 Complete the equation in your answer booklet for the neutralization that occurs in the first titration by writing a formula of the missing product. [1]



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