

Quizzes from

Jan. 2010

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Table A
Standard Temperature and Pressure

Measure	Value	Unit
Standard Pressure	101.3 kPa 1 atm	kilopascals atmospheres
Standard Temperature	273 K 0 °C	kelvin degrees Celsius

- 7) Which sample of matter has the greatest increase between its boiling and STP?

- (1) H_2O (2) C_2H_6 (3) C_2H_4 (4) C_2H_2

The melting points and boiling points of five substances at standard pressure are listed in the table below.

Melting Points and Boiling Points of Five Substances

Substance	Melting Point (°C)	Boiling Point (°C)
HCl	-109	-85
H_2O	0	100
F_2	-93	-84
Br_2	-7.3	58.8
I_2	113.7	184.3

- 7) Which substance in the table has a liquid at STP? (1)



- 8) Convert the liquid at standard pressure of the substance to atmospheres.

Regents Chemistry New York Lab Manual

**Charles Alpheus Bennett, William
Thomas Bawden**



Regents Chemistry New York Lab Manual:

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause-effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

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