Redox Stoichiometry (Titrations)

- In a redox titration, a known concentration of an oxidizing agent is used to find an unknown concentration of a reducing agent, or vice-versa.
- Applications include finding the iron content of water or the Vitamin C content in foods.

General Steps:

- 1. Balance the redox equation
- 2. Organize data for the given and required substances
- Find the moles of the given substance

If given volume and concentration, use n=CvIf given mass, find molar mass, then use n=mM

- 4. Find the moles of the required substances, using mole ratio
- Convert moles to concentration (C=n/v) or volume (v=n/C) or mass (m=nM) of required substance

Example 1:

The following example involves the titration of a solution of sodium oxalate, $Na_2C_2O_{4(aq)}$ with a solution of potassium permanganate, $KMnO_{4(aq)}$.

Important Note:

- The permanganate ion, MnO₄, is often used as an oxidizing agent because of its deep purple color.
- It acts as its own indicator.
- It is reduced to pale pink Mn2+ by the oxalate ions.
- Once all oxalate ions have been oxidized the next drop of permanganate ions turns the solution a faint purple color, signaling the endpoint of the titration.

Find the **volume** of 0.200 M solution of MnO_4 which will react with 50.0 mL of 0.100 M solution of $C_2O_4^{2}$.

$$MnO_4^- + C_2O_4^{2-} \rightarrow Mn^{2+} + CO_2$$

Example 2:

A 23.30 ml sample of KMnO₄ solution is decolorized by 0.1111g of oxalic acid $(H_2C_2O_4)$. The products are Mn^{2+} and CO_2 gas. Calculate the concentration of KMnO₄ using the following equation:

Example 3:

The oxidation of Fe²⁺ ions by permanganate in acidic solution is accompanied by the formation of Fe³⁺ ions and Mn²⁺ ions. How many moles of FeSO₄ would be oxidized by 100.0 ml of 0.02118 mol/L KMnO₄ solution.

$$Fe^{2+} + MnO_4 \rightarrow Fe^{3+} + Mn^{2+}$$

Redox Stoichiometry Worksheet

Malika Ammam

Redox Stoichiometry Worksheet:

Basics of Analytical Chemistry and Chemical Equilibria Brian M. Tissue, 2013-06-06 Enables students to progressively build and apply new skills and knowledge Designed to be completed in one semester this text enables students to fully grasp and apply the core concepts of analytical chemistry and aqueous chemical equilibria Moreover the text enables readers to master common instrumental methods to perform a broad range of quantitative analyses Author Brian Tissue has written and structured the text so that readers progressively build their knowledge beginning with the most fundamental concepts and then continually applying these concepts as they advance to more sophisticated theories and applications Basics of Analytical Chemistry and Chemical Equilibria is clearly written and easy to follow with plenty of examples to help readers better understand both concepts and applications In addition there are several pedagogical features that enhance the learning experience including Emphasis on correct IUPAC terminology You Try It spreadsheets throughout the text challenging readers to apply their newfound knowledge and skills Online tutorials to build readers skills and assist them in working with the text s spreadsheets Links to analytical methods and instrument suppliers Figures illustrating principles of analytical chemistry and chemical equilibria End of chapter exercises Basics of Analytical Chemistry and Chemical Equilibria is written for undergraduate students who have completed a basic course in general chemistry In addition to chemistry students this text provides an essential foundation in analytical chemistry needed by students and practitioners in biochemistry environmental science chemical engineering materials science nutrition agriculture and the life sciences

General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material Chemistry, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics. The series will be aligned to current science standards **Chemistry** Carson-Dellosa Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements

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increasingly becoming a consideration in synthesis and functional chemistry and an essential element for the next generation of chemistry in academia and industry This book represents a compilation of the latest advancements in functional redox chemistry and demonstrates its importance in achieving a more sustainable future This book is an ideal companion for any postgraduate students or researchers interested in sustainability in academia and industry

Redox Models in Chemistry Lise-Lotte Österlund, Margareta Ekborg, Anders Berg, 2010

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