

# BALANCING REDOX REACTIONS

STEP BY STEP IN TWO WAYS...



## BOTH START THE SAME WAY

1. ASSIGN OXIDATION NUMBERS TO EACH ATOM



NOTE: OXIDATION NUMBERS SUM UP TO ZERO EVERYWHERE

ALSO, THERE IS ALWAYS AN OXIDATION STATE CHANGE

IF SOMETHING IS NOT AN OXIDATION STATE CHANGE

2. WRITE YOUR EQUATION AS TWO HALF REACTIONS (BALANCING ALL ATOMS EXCEPT H AND O)

NOTE: ONLY H<sup>+</sup> & OH<sup>-</sup> CAN BE USED TO BALANCE THESE



## HALF-REACTION METHOD (in acidic conditions)

3. BALANCE O WITH H<sub>2</sub>O



4. BALANCE H WITH H<sup>+</sup>



5. BALANCE CHARGES WITH ELECTRONS

NOTE: ELECTRONS MUST ALWAYS BE CONSUMED & NOT PRODUCED IN THE END



6. MAKE SURE THE NUMBER OF ELECTRONS IS THE SAME

NOTE: YOU CAN BALANCE WITH AN ELECTRONS IN THE EQUATION (NOT RECOMMENDED)



7. PUT BOTH REACTIONS BACK TOGETHER



8. CHECK FOR ANYTHING THAT IS BOTH A PRODUCT & REACTANT

NOTE: SOME REACTIONS CAN BE DONE WITH OH<sup>-</sup> INSTEAD OF H<sup>+</sup>



## BUT WHAT ABOUT BASIC CONDITIONS?

WHEN WE SAY "ACIDIC" OR "BASIC" WE ARE USING TWO APPROXIMATE DEFINITIONS:

ACIDIC CONDITIONS HAVE H<sup>+</sup> AND BASIC CONDITIONS HAVE OH<sup>-</sup>

SO WHEN WE FIND OUR REACTION FROM ABOVE, WITH A REACTION PERFORMED IN BASIC CONDITIONS:

8. ADD THE NUMBER OF OH<sup>-</sup> TO THE NUMBER OF H<sup>+</sup> IN EACH SIDE

NOTE: ONLY ONE OH<sup>-</sup> CAN BE USED ON EACH SIDE (NOT BOTH)



# Redox Reactions Teacher Guide And Answers

**Lauren Gardner**



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