

Worksheet 25 - Oxidation/Reduction Reactions

Oxidation number rules:

Elements have an oxidation number of **0**

Group I and II – In addition to the elemental oxidation state of 0, Group I has an oxidation state of **+1** and Group II has an oxidation state of **+2**.

Hydrogen – usually **+1**, except when bonded to Group I or Group II, when it forms hydrides, **-1**.

Oxygen – usually **-2**, except when it forms a O-O single bond, a peroxide, when it is **-1**.

Fluorine is always **-1**. Other halogens are usually **-1**, except when bonded to O.

1. Assign **oxidation numbers** to each of the atoms in the following compounds:



- a. What is the range of oxidation states for **carbon**? **-4 to +4**
- b. Which compound has C in a **+4** state? **CO_2**
- c. Which compound has C in a **-4** state? **CH_4**

Redox Ws 2 Oxidation Numbers Answers

Rachel S Tattersall



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