

Redox practice worksheet

Name: _____

Date: _____

- In which substance is the oxidation number of nitrogen zero?
A. NH_3 B. N_2 C. NO_2 D. N_2O
- What is the oxidation number of carbon in NaHCO_3 ?
A. +6 B. +2 C. -4 D. +4
- In the reaction $\text{Al}^0 + \text{Cr}^{3+} \rightarrow \text{Al}^{3+} + \text{Cr}^0$, the reducing agent is
A. Al^0 B. Cr^{3+} C. Al^{3+} D. Cr^0

- In the reaction $2\text{K} + \text{Cl}_2 \rightarrow 2\text{KCl}$, the species oxidized is
A. Cl_2 B. Cl^- C. K D. K^+
- As an S^{2-} ion is oxidized to an S^0 atom, the number of protons in its nucleus
A. decreases B. increases
C. remains the same

- Given the probable reaction for the nickel-cadmium battery:
$$2\text{Ni}(\text{OH})_2 + \text{Cd} \rightleftharpoons 2\text{Ni}(\text{OH})_2 + \text{Cd}(\text{OH})_2$$

Which species is oxidized during the discharge of the battery?

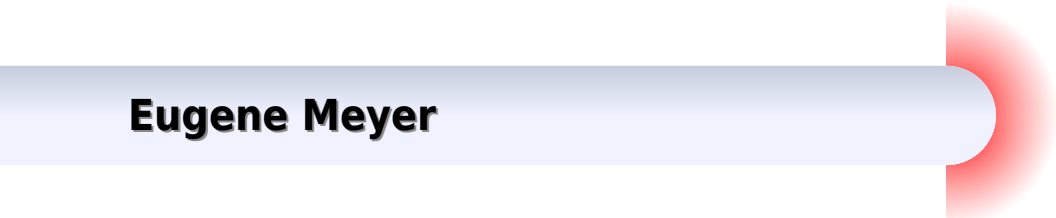
A. Ni^{2+} B. Ni^{2+} C. Cd^0 D. Cd^{2+}

- If element X forms the oxides XO and X_2O_3 , the oxidation numbers of element X are
A. +1 and +2 B. +2 and +3
C. +1 and +3 D. +2 and +4
- Oxygen has a positive oxidation number in the compound
A. H_2O B. H_2O_2 C. OF_2 D. IO_2
- What is the oxidation number of sulfur in H_2SO_4 ?
A. 0 B. -2 C. +6 D. +4

- In the equation $\text{Cu(s)} + 2\text{Ag}^+(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq}) + 2\text{Ag(s)}$, the oxidizing agent is
A. Cu^0 B. Ag^+ C. Cu^{2+} D. Ag^0
- In the reaction $\text{Al} + \text{Cr}^{3+} \rightarrow \text{Al}^{3+} + \text{Cr}$, the reducing agent is
A. Al B. Cr^{3+} C. Al^{3+} D. Cr
- In the compound Na_2HPO_4 , which element has a negative oxidation number?
A. H B. O C. P D. Na

Redox Reactions Answers Holt Chemistry

Eugene Meyer



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Holt Chemistry Ralph Thomas Myers, 2004 **Holt Chemistry** Holt Rinehart & Winston, 2003-01-24 National 5 Chemistry with Answers, Second Edition Barry McBride, Stephen Jeffrey, John Anderson, Paul McCranor, Fran Macdonald, 2018-07-30 Exam Board SQA Level National 5 Subject Chemistry First Teaching September 2017 First Exam Summer 2018 The second edition of this textbook has been fully revised and updated to reflect changes made to the SQA syllabus from 2017 onwards New features include Refreshed content Additional candidate advice Model answers for open ended questions *Holt Chemistry File*, 1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket **Fundamentals of Chemistry: A Modern Introduction (1966)** Frank Brescia, 2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro's Law and Graham's Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry *Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Redox Chemistry and Interfacial Behavior of Biological Molecules Glenn Dryhurst, K Niki, 2012-12-06 The papers in this book were presented at the Third International Symposium on Redox Mechanisms and Interfacial Properties of Molecules of Biological Importance held in Honolulu Hawaii between October 19-23 1987 This Symposium was held as part of the 172nd Meeting of The Electrochemical Society which was cosponsored by The Electrochemical Society of Japan with the cooperation of The Japan Society of Applied Physics The aim of the Symposium was to bring together a group of electrochemists and bio medical scientists with interests in electrochemistry from around the world to present their most current research results and or to present up to date reviews of current areas of research activity It is quite clear from the diversity of topics covered in the various papers that electrochemistry and electrochemical techniques and principles have much to contribute to our under

standing of many important biochemical phenomena For example electrochemical studies are providing important insights into the redox properties of biomolecules ranging from relatively small organic molecules such as indoleamine neurotransmitters to very large organic organometallic molecules which include various redox enzymes or model enzyme systems Many of the most powerful analytical techniques are now being coupled to electrodes to monitor potential controlled behaviors of biological molecules at charged interfaces Electrochemical techniques are now being developed which permit extraordinarily small electrodes to be inserted into single cells to monitor electroactive biomolecules Other microelectrodes are being employed to control cell growth and to manipulate single cells

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Introduction to Modern Chemistry: Student Manual Eugene Meyer, 1979 **Fundamentals of Chemistry: A Modern Introduction** Frank Brescia, 2012-12-02

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