

Redox/Electrochemistry Regents Unit Review

Part I:

1. Base your answers on the following redox reaction, which occurs spontaneously in an electrochemical cell:



- Write the half-reaction for the reduction that occurs.
- Write the half-reaction for the oxidation that occurs.
- Balance the equation using the smallest whole-number coefficients.
- Which species loses electrons and which species gains electrons?
- Which half-reaction occurs at the cathode?
- State what happens to the number of protons in a Zn atom when it changes to Zn^{2+} as the redox reaction occurs.

2. Given the reaction:



Which species undergoes oxidation?

- (1) Mg(s) (2) $\text{H}^+(\text{aq})$ (3) $\text{Cl}^-(\text{aq})$ (4) $\text{H}_2(\text{g})$
3. Which particles are gained and lost during a redox reaction?
- (1) electrons (2) protons (3) neutrons (4) positrons
4. As a Ca atom undergoes oxidation to Ca^{2+} , the number of neutrons in its nucleus
- (1) decreases (2) increases (3) remains the same

5. Given the reaction:



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