

Name: \_\_\_\_\_  
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Date: \_\_\_\_\_



**Exceed to Succeed**

## **Worksheet: Periodic Trends**

### **1. ATOMIC RADIUS**

For each of the following sets of atoms, rank the atoms from smallest to largest atomic radius.

- a. Li, C, F
- b. Li, Na, K
- c. Ge, P, O
- d. C, N, Al
- e. Al, Cl, Ga

### **2. IONIC RADIUS**

For each of the following sets of ions, rank them from smallest to largest ionic radius.

- a.  $\text{Mg}^{2+}$ ,  $\text{Si}^{4+}$ ,  $\text{S}^{2-}$
- b.  $\text{Mg}^{2+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Ba}^{2+}$
- c.  $\text{F}^-$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$
- d.  $\text{Ba}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Zn}^{2+}$
- e.  $\text{Si}^{4+}$ ,  $\text{P}^{3-}$ ,  $\text{O}^{2-}$

### **3. IONIZATION ENERGY**

For each of the following sets of atoms, rank them from lowest to highest ionization energy.

- a. Mg, Si, S
- b. Mg, Ca, Ba
- c. F, Cl, Br
- d. Ba, Cu, Ne
- e. Si, P, He

### **4. ELECTRONEGATIVITY**

For each of the following sets of atoms, rank them from lowest to highest electronegativity.

- a. Li, C, N
- b. C, O, Ne
- c. Si, P, O
- d. K, Mg, P
- e. S, F, He

# Periodic Trends And Properties Of Elements Lab Answer Keys

**Jo Allan Beran**



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