

# Graphing Periodic Trends Activity

Name: \_\_\_\_\_ Period: \_\_\_\_\_

Use the information in these tables to complete the graph as described below

| Symbol | Atomic Radius (Picometers) | First Ionization energy (kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|-------------------------------------------|-----------------------------------|
| H      | 31                         | 1312                                      | 2.1                               |
| Li     | 128                        | 520                                       | 1.0                               |
| Na     | 166                        | 496                                       | 0.9                               |
| K      | 203                        | 410                                       | 0.8                               |
| Rb     | 220                        | 403                                       | 0.8                               |
| Cs     | 244                        | 376                                       | 0.7                               |

| Symbol | Atomic Radius (Picometers) | First Ionization energy (kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|-------------------------------------------|-----------------------------------|
| Na     | 166                        | 496                                       | 0.9                               |
| Mg     | 141                        | 738                                       | 1.2                               |
| Al     | 121                        | 578                                       | 1.5                               |
| Si     | 111                        | 787                                       | 1.8                               |
| P      | 107                        | 1012                                      | 2.1                               |
| S      | 105                        | 1000                                      | 2.5                               |
| Cl     | 102                        | 1251                                      | 3.0                               |
| Ar     | 106                        | 1521                                      | ---                               |

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

|                         |                                                                                                                                                                                          |    |    |   |   |    |    |
|-------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----|----|---|---|----|----|
| H                       | <p>Across a period (→) atomic radius tends to _____.</p> <p>Across a period (→) first ionization energy tends to _____.</p> <p>Across a period (→) electronegativity tends to _____.</p> |    |    |   |   |    |    |
| Li                      |                                                                                                        |    |    |   |   |    |    |
| Na<br>166<br>496<br>0.9 | Mg                                                                                                                                                                                       | Al | Si | P | S | Cl | Ar |
| K                       | <p>Down a Group (↓) atomic radius tends to _____.</p> <p>Down a Group (↓) first ionization energy tends to _____.</p> <p>Down a Group (↓) electronegativity tends to _____.</p>          |    |    |   |   |    |    |
| Rb                      |                                                                                                                                                                                          |    |    |   |   |    |    |
| Cs                      |                                                                                                                                                                                          |    |    |   |   |    |    |

# Regular Chemistry Periodic Trends Activity Answer Key

**The Princeton Review**



## **Regular Chemistry Periodic Trends Activity Answer Key:**

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### **Table of Contents Regular Chemistry Periodic Trends Activity Answer Key**

1. Understanding the eBook Regular Chemistry Periodic Trends Activity Answer Key
  - The Rise of Digital Reading Regular Chemistry Periodic Trends Activity Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Regular Chemistry Periodic Trends Activity Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Regular Chemistry Periodic Trends Activity Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Regular Chemistry Periodic Trends Activity Answer Key
  - Personalized Recommendations
  - Regular Chemistry Periodic Trends Activity Answer Key User Reviews and Ratings

- Regular Chemistry Periodic Trends Activity Answer Key and Bestseller Lists
- 5. Accessing Regular Chemistry Periodic Trends Activity Answer Key Free and Paid eBooks
  - Regular Chemistry Periodic Trends Activity Answer Key Public Domain eBooks
  - Regular Chemistry Periodic Trends Activity Answer Key eBook Subscription Services
  - Regular Chemistry Periodic Trends Activity Answer Key Budget-Friendly Options
- 6. Navigating Regular Chemistry Periodic Trends Activity Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Regular Chemistry Periodic Trends Activity Answer Key Compatibility with Devices
  - Regular Chemistry Periodic Trends Activity Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Regular Chemistry Periodic Trends Activity Answer Key
  - Highlighting and Note-Taking Regular Chemistry Periodic Trends Activity Answer Key
  - Interactive Elements Regular Chemistry Periodic Trends Activity Answer Key
- 8. Staying Engaged with Regular Chemistry Periodic Trends Activity Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Regular Chemistry Periodic Trends Activity Answer Key
- 9. Balancing eBooks and Physical Books Regular Chemistry Periodic Trends Activity Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Regular Chemistry Periodic Trends Activity Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Regular Chemistry Periodic Trends Activity Answer Key
  - Setting Reading Goals Regular Chemistry Periodic Trends Activity Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Regular Chemistry Periodic Trends Activity Answer Key
  - Fact-Checking eBook Content of Regular Chemistry Periodic Trends Activity Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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