

Electrochemistry Regents Questions

Chemistry Regents June 2008

26 Which half-reaction equation represents the reduction of a potassium ion?



A flashlight can be powered by a rechargeable nickel-cadmium battery. In the battery, the anode is $\text{Cd}(s)$ and the cathode is $\text{NiO}_2(s)$. The unbalanced equation below represents the reaction that occurs as the battery produces electricity. When a nickel-cadmium battery is recharged, the reverse reaction occurs.



75 Determine the change in oxidation number for the element that makes up the anode in the reaction that produces electricity. [1]

From _____ to _____

76 Explain why Cd would be above Ni if placed on Table J. [1]

51. Base your answer on the information below.

The unbalanced equation below represents the decomposition of potassium chlorate.



a Balance the equation in the space provided or on a separate piece of paper, using the smallest whole-number coefficients. [1]

b Determine the oxidation number of chlorine in the reactant. [1]



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