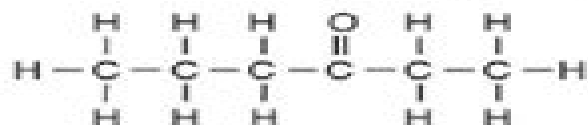


57. Given the formula representing a compound:



What is an IUPAC name for this compound?

- A. ethyl propanoate
  - B. propyl ethanoate
  - C. 3-hexanone
  - D. 4-hexanone
58. As the temperature of a reaction increases, it is expected that the reacting particles collide
- A. more often and with greater force
  - B. more often and with less force
  - C. less often and with greater force
  - D. less often and with less force
59. What is the total amount of heat required to completely melt 347 grams of ice at its melting point?
- A. 334 J
  - B. 1450 J
  - C. 116,000 J
  - D. 784,000 J
60. Which two notations represent isotopes of the same element?
- A.  ${}^9_7\text{N}$  and  ${}^{10}_7\text{N}$
  - B.  ${}^{20}_7\text{N}$  and  ${}^{20}_{10}\text{Ne}$
  - C.  ${}^9_7\text{N}$  and  ${}^{10}_{10}\text{Ne}$
  - D.  ${}^{10}_7\text{N}$  and  ${}^{10}_{10}\text{Ne}$
61. Which pair of atoms has the most polar bond?
- A. H-Br
  - B. H-Cl
  - C. I-Br
  - D. I-Cl
62. What occurs when potassium reacts with chlorine to form potassium chloride?
- A. Electrons are shared and the bonding is ionic.
  - B. Electrons are shared and the bonding is covalent.
  - C. Electrons are transferred and the bonding is ionic.
  - D. Electrons are transferred and the bonding is covalent.

63. Given the balanced equation representing a reaction:



What occurs as bonds are broken in one mole of  $\text{H}_2$  molecules during this reaction?

- A. Energy is absorbed and one mole of unbonded hydrogen atoms is produced.
  - B. Energy is absorbed and two moles of unbonded hydrogen atoms are produced.
  - C. Energy is released and one mole of unbonded hydrogen atoms is produced.
  - D. Energy is released and two moles of unbonded hydrogen atoms are produced.
64. Which equation shows conservation of charge?
- A.  $\text{Cu} + \text{Ag}^+ \rightarrow \text{Cu}^{2+} + \text{Ag}$
  - B.  $\text{Mg} + \text{Zn}^{2+} \rightarrow 2\text{Mg}^{2+} + \text{Zn}$
  - C.  $2\text{F}_2 + \text{Br}^- \rightarrow 2\text{F}^- + \text{Br}_2$
  - D.  $2\text{I}^- + \text{Cl}_2 \rightarrow \text{I}_2 + 2\text{Cl}^-$
65. At STP, graphite and diamond are two solid forms of carbon. Which statement explains why these two forms of carbon differ in hardness?
- A. Graphite and diamond have different ionic radii.
  - B. Graphite and diamond have different molecular structures.
  - C. Graphite is a metal, but diamond is a nonmetal.
  - D. Graphite is a good conductor of electricity, but diamond is a poor conductor of electricity
66. Which electron configuration represents an excited state for an atom of calcium?
- A. 2-8-7-1
  - B. 2-8-7-2
  - C. 2-8-7-3
  - D. 2-8-8-2
67. Which list of elements is arranged in order of increasing electronegativity?
- A. Be, Mg, Ca
  - B. F, Cl, Br
  - C. K, Ca, Sc
  - D. Li, Na, K
68. Which radioisotope has the fastest rate of decay?
- A.  ${}^{14}\text{C}$
  - B.  ${}^{40}\text{Ca}$
  - C.  ${}^{235}\text{Pu}$
  - D.  ${}^{40}\text{K}$

# Regents Chem The Periodic Law Worksheet

**Albert Tarendash**



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