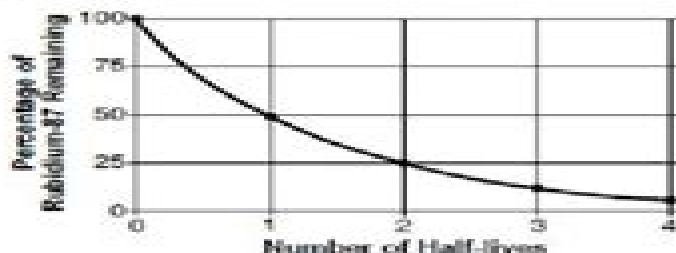


Name: _____

Answer all questions on the answer document. This is a TAKE HOME TEST due on Wednesday, June 7.

1. The graph below shows the radioactive decay of rubidium-87.



What percentage of rubidium-87 atoms will be left after four half-lives?

- A. 25.0%
B. 12.5%
C. 6.25%
D. 3.125%
2. Given the equation representing a reaction at equilibrium:
- $$\text{H}_2\text{S}(\text{aq}) + \text{CH}_3\text{NH}_2(\text{aq}) \rightleftharpoons \text{HS}^-(\text{aq}) + \text{CH}_3\text{NH}_3^+(\text{aq})$$
- According to one acid-base theory, the forward reaction is classified as an acid-base reaction because
- A. H_2S is a H^+ donor and CH_3NH_2 is a H^+ acceptor
B. CH_3NH_2 is a H^+ donor and H_2S is a H^+ acceptor
C. HS^- and CH_3NH_3^+ are both H^+ donors
D. CH_3NH_3^+ and HS^- are both H^+ acceptors
3. Which ionic equation represents a spontaneous reaction that can occur in a voltaic cell?
- A. $\text{Cu}(\text{s}) + \text{Zn}(\text{s}) \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Zn}^{2+}(\text{aq})$
B. $\text{Cu}(\text{s}) + \text{Zn}^{2+}(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Zn}(\text{s})$
C. $\text{Cu}^{2+}(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{Cu}(\text{s}) + \text{Zn}^{2+}(\text{aq})$
D. $\text{Cu}^{2+}(\text{aq}) + \text{Zn}^{2+}(\text{aq}) \rightarrow \text{Cu}(\text{s}) + \text{Zn}(\text{s})$

4. Given the equation representing a reaction:
- $$\text{Sn}^{4+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn}^{2+}(\text{aq})$$

Which term best describes this reaction?

- A. ionization
B. neutralization
C. oxidation
D. Reduction

5. The table below shows the atomic mass and natural abundance of the two naturally occurring isotopes of lithium.

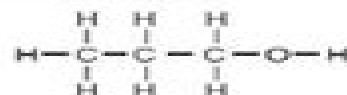
Naturally Occurring Isotopes of Lithium

Isotope	Atomic Mass (u)	Natural Abundance (%)
Li-6	6.015	7.6
Li-7	7.016	92.4

Which numerical setup can be used to determine the atomic mass of naturally occurring lithium?

- A. $(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})$
B. $(0.076)(6.015 \text{ u}) + (0.924)(7.016 \text{ u})$
C. $\frac{(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})}{2}$
D. $\frac{(0.076)(6.015 \text{ u}) + (0.924)(7.016 \text{ u})}{2}$

6. Given the formulas representing two compounds at standard pressure:



1-propanol



methyl ethyl ether

The compounds can be differentiated by their

- A. boiling points
B. gram-formula masses
C. numbers of hydrogen atoms
D. percent compositions by mass of carbon
7. Given the formula representing a compound:



What is the IUPAC name of this compound?

- A. 2-chloroheptane
B. 6-chloroheptane
C. 2,2-dichloroheptane
D. 6,6-dichloroheptane

Regents Chemistry The Periodic Law Worksheet

Michael Rajnik



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