

The Periodic Table of Elements

In 1869 Dimitri Mendeleev organized the elements in a table based on the properties of the elements.

Groups

↓ ↓
(Columns)

The elements in each group have the same number of valence electrons.

1 st Period → (1 shell)	1 H Hydrogen	Group 2-8 (8 valence electrons)	Group 3-10 (8 valence electrons)	Group 11 (1 valence electron)	Group 12 (2 valence electrons)	Group 13 (3 valence electrons)	Group 14 (4 valence electrons)	Group 15 (5 valence electrons)	Group 16 (6 valence electrons)	Group 17 (7 valence electrons)	Group 18 (8 valence electrons)	2 He Helium	↳ Except Helium with 2 valence e ⁻
2 nd Period → (2 shells)	3 Li Lithium	4 Be Beryllium	5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon					
3 rd Period → (3 shells)	11 Na Sodium	12 Mg Magnesium	13 Al Aluminum	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon					
4 th Period → (4 shells)	19 K Potassium	20 Ca Calcium	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton					

Periods
(rows)

The elements in each period have the same number of shells.

Today, 114 elements are arranged by atomic number in the periodic table. The atomic number is based on the number of protons in an element.

Periodic Table Notes For Middle School

Sebastian Brünink



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Table of Contents Periodic Table Notes For Middle School

1. Understanding the eBook Periodic Table Notes For Middle School
 - The Rise of Digital Reading Periodic Table Notes For Middle School
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Notes For Middle School
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Notes For Middle School
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Notes For Middle School
 - Personalized Recommendations
 - Periodic Table Notes For Middle School User Reviews and Ratings
 - Periodic Table Notes For Middle School and Bestseller Lists
5. Accessing Periodic Table Notes For Middle School Free and Paid eBooks
 - Periodic Table Notes For Middle School Public Domain eBooks
 - Periodic Table Notes For Middle School eBook Subscription Services
 - Periodic Table Notes For Middle School Budget-Friendly Options

6. Navigating Periodic Table Notes For Middle School eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Notes For Middle School Compatibility with Devices
 - Periodic Table Notes For Middle School Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Notes For Middle School
 - Highlighting and Note-Taking Periodic Table Notes For Middle School
 - Interactive Elements Periodic Table Notes For Middle School
8. Staying Engaged with Periodic Table Notes For Middle School
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Notes For Middle School
9. Balancing eBooks and Physical Books Periodic Table Notes For Middle School
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Notes For Middle School
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Notes For Middle School
 - Setting Reading Goals Periodic Table Notes For Middle School
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Notes For Middle School
 - Fact-Checking eBook Content of Periodic Table Notes For Middle School
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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