

Annotated periodic table of the elements

1	H 1.008 Hydrogen																		18	Ar 39.948 Argon			
2	Li 6.94 Lithium		Be 9.012 Beryllium																		19	K 39.098 Potassium	
3	Na 22.99 Sodium		Mg 24.305 Magnesium																		20	Ca 40.078 Calcium	
4	K 39.098 Potassium		Ca 40.078 Calcium																		21	Sc 44.956 Scandium	
5	Rb 85.468 Rubidium		Sr 87.62 Strontium																		22	Zr 91.224 Zirconium	
6	Cs 132.905 Cesium		Ba 137.327 Barium																		23	Nb 92.906 Niobium	
7	Fr 223.019 Francium		Ra 226.025 Radium																		24	Mo 95.94 Molybdenum	
8																					25	Tc 98.906 Technetium	
9																					26	Ru 101.07 Ruthenium	
10																					27	Rh 102.905 Rhodium	
11																					28	Pd 106.42 Palladium	
12																					29	Ag 107.868 Silver	
13																					30	Cd 112.411 Cadmium	
14																					31	In 114.818 Indium	
15																					32	Sn 118.710 Tin	
16																					33	Sb 121.757 Antimony	
17																					34	Te 127.60 Tellurium	
18																					35	I 126.905 Iodine	
19																					36	Xe 131.29 Xenon	
20																					37	Ba 137.327 Barium	
21																					38	La 138.905 Lanthanum	
22																					39	Ce 140.12 Cerium	
23																					40	Pr 140.908 Praseodymium	
24																					41	Nd 144.24 Neodymium	
25																					42	Pm 144.913 Promethium	
26																					43	Sm 150.36 Samarium	
27																					44	Eu 151.964 Europium	
28																					45	Gd 157.25 Gadolinium	
29																					46	Tb 158.925 Terbium	
30																					47	Dy 162.50 Dysprosium	
31																					48	Ho 164.930 Holmium	
32																					49	Er 167.26 Erbium	
33																					50	Tm 168.933 Thulium	
34																					51	Yb 173.054 Ytterbium	
35																					52	Lu 174.967 Lutetium	
36																					53	Hf 178.49 Hafnium	
37																					54	Ta 180.948 Tantalum	
38																					55	W 183.84 Tungsten	
39																					56	Re 186.207 Rhenium	
40																					57	Os 190.23 Osmium	
41																					58	Ir 192.22 Iridium	
42																					59	Pt 195.084 Platinum	
43																					60	Au 196.967 Gold	
44																					61	Hg 200.59 Mercury	
45																					62	Tl 204.38 Thallium	
46																					63	Pb 207.2 Lead	
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48																					65	Po 209 Polonium	
49																					66	At 210 Astatine	
50																					67	Rn 222 Radon	
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CLASS COUNTERPARTS		
Metals		Nonmetals
Active	Pre-transition	Comet
	Apex earth	
Transition		Intermediate
Frontier		Metalloid
Noble (Ru-Pt, Os-Au)		Noble gas

57 La 138.91 Lanthanum	58 Ce 140.12 Cerium	59 Pr 140.91 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 144.91 Promethium	62 Sm 150.36 Samarium	63 Eu 151.96 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.93 Terbium	66 Dy 162.50 Dysprosium	67 Ho 164.93 Holmium	68 Er 167.26 Erbium	69 Tm 168.93 Thulium	70 Yb 173.05 Ytterbium
89 Ac 227.03 Actinium	90 Th 232.04 Thorium	91 Pa 231.04 Protactinium	92 U 238.03 Uranium	93 Np 237.05 Neptunium	94 Pu 244.06 Plutonium	95 Am 243.06 Americium	96 Cm 247.07 Curium	97 Bk 247.07 Berkelium	98 Cf 251.08 Californium	99 Es 252.08 Einsteinium	100 Fm 257.10 Fermium	101 Md 258.10 Mendelevium	102 No 259.10 Nobelium

Active metals

Pre-transition
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zirconates, titanates, and tantalates.

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750-1,000 kJ/mol), and electronegativity values (1.8-2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colourless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

Group 1
This group is Sc-Y, and the alkaline earths are La-Be, and 4 to 10 (not 11-12). Trends going down Sc-Y and Be:

- La trend, resemble those of groups 1-2
- Lu trend, more resemble those of groups 4-10

Comet nonmetals
O, F, and Cl are, respectively, colourless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically comet nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colourless, odourless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answer Sheet

Huangqi Zhang

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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