

**Multiple Choice:** *Identify the letter of the choice that best completes the statement or answers the question.*

- \_\_\_\_ 1. What scientific (natural) law serves as the basis for balancing chemical equations by requiring that there be no observable change in the quantity of matter during a chemical reaction?
- Law of Conservation of Energy
  - Law of Multiple Proportions
  - Law of Conservation of Matter
  - Law of Definite Proportions
  - Law of Constant Composition
- \_\_\_\_ 2. Balance the following equation with the **smallest whole number coefficients**. What is the coefficient for  $O_2$  in the balanced equation?



- 9
  - 5
  - 15
  - 6
  - 13
- \_\_\_\_ 3. What is the coefficient for HBr when the following equation is balanced with the **smallest whole number coefficients**?
- $$Br_2 + H_2O \rightarrow HBr + HBrO_3$$
- 5
  - 7
  - 8
  - 3
  - 6
- \_\_\_\_ 4. Balance the following equation with the **smallest whole number coefficients**. What is the coefficient for  $NH_3$  in the balanced equation?
- $$Fe(NO_3)_3 + NH_3 + H_2O \rightarrow Fe(OH)_3 + NH_4NO_3$$
- 1
  - 3
  - 2
  - 6
  - 4
- \_\_\_\_ 5. Elemental phosphorus is produced from calcium phosphate in the following reaction. What is the coefficient for C when this equation is balanced with the **smallest whole number coefficients**?
- $$Ca_3(PO_4)_2 + SiO_2 + C \rightarrow P_4 + CO + CaSiO_3$$

- 10
- 3
- 1
- 6
- 4

# Plato Earth Science Semester 1 Answers

**Rachel S Tattersall**



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