

1. Particles

- If we could divide any substance down into smaller and smaller pieces we could see it is made of tiny particles.
- These particles are so small they could not be seen using a microscope.
- We can represent these particles in models using spheres.



4. Changes in states of matter



- A solid melts when it is heated because the particles are **gaining energy**. This energy is used to break the forces of attraction between the molecules. The more energy a particle has the **faster** it can move.

6. Elements and mixtures

- This diagram shows a mixture of elements.
- Mixtures can be easily separated using techniques such as:
 - Filtration
 - Distillation
 - Chromatography



2. States of matter

- Three states of matter are solid, liquid and gas.

Solid	Liquid	Gas

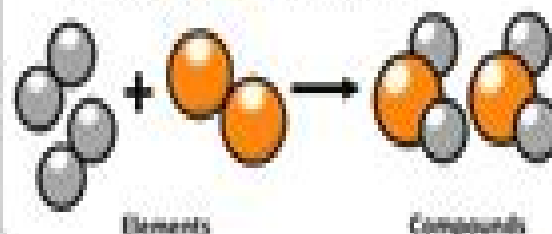


KS3 Spine

All matter in the universe
is made of small particles

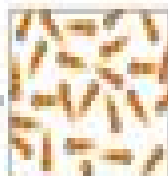
7. Compounds

- **Atoms** of different **elements** can combine together to form a very large number of **compounds**.



3. Differences in states of matter

- In solids the particles are **closely spaced and vibrating**. In liquids there is **random motion** but the particles are in **contact**.
- In a gas there is **random motion** and the particles are **widely spaced**.
- Liquids and gases **flow** and completely fill their container. Gases can be **compressed**.



5. Atoms and elements

- **Atoms** are the **basic building blocks** of all living and non-living things anywhere in the universe.
- There are just over 100 different types of **atoms**.
- Substances made of only **one kind of atom** are called **elements**.
- All the known elements are listed in the **periodic table of elements**.



8. Reactions and properties

- A **chemical reaction** involves the **rearrangement** of atoms to form new substances, while the **total number of atoms stays the same** (see image above).
- The **properties** of different materials can be explained in terms of the **behaviour of the atoms and groups of atoms** of which they are made.

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