

$$M = \frac{\# \text{ of moles}}{V \text{ (in L)}}$$

Molar Concentration

Part A: Describing Molar Concentration

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Explain what this concentration tells us about the NaCl solution. How might a chemist use this ratio?

$$\frac{1.5 \text{ mol of NaCl}}{1 \text{ L of water}}$$

if you need a certain amount of NaCl you can take a certain volume of the solution.

2. A 45.3 g sample of potassium nitrate is dissolved in enough water to make 225 mL of solution. Determine the molar concentration of the potassium nitrate.

$$\begin{aligned} \text{KNO}_3 : \text{m.m.} &= 101.11 \text{ g/mol} & 0.45 \text{ mol KNO}_3 \\ V &= 0.225 \text{ L} & \frac{0.45 \text{ mol KNO}_3}{0.225 \text{ L}} = 2 \text{ M} \end{aligned}$$

3. Find the molarity of a solution made from 275 g of CuSO_4 dissolved in enough water to make 4.25 L.

$$\begin{aligned} 275 \text{ g CuSO}_4 & \quad \frac{275 \text{ g}}{159.65} = 1.72 \text{ mol CuSO}_4 \\ \text{m.m.} & \quad \frac{1.72 \text{ mol}}{4.25 \text{ L}} = 0.40 \text{ M CuSO}_4 \\ 63.55 & \\ + 32.1 & \\ \hline 95.65 & \end{aligned}$$

4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.

$$\begin{aligned} 5.15 \text{ g I}_2 & \quad \frac{5.15 \text{ g I}_2}{253.8} = 0.02 \text{ mol I}_2 \\ 253.8 & \quad \frac{0.02 \text{ mol I}_2}{0.225 \text{ L}} = 0.089 \text{ M I}_2 \end{aligned}$$

5. What final volume would be needed in order to prepare a 0.25 M NaCl solution from 5.2 g of NaCl (s)?

$$\begin{aligned} \text{m.m. NaCl} : & \quad \frac{5.2 \text{ g}}{58.44 \text{ g/mol}} = 0.089 \text{ mol NaCl} & M = \frac{\# \text{ of mol}}{V \text{ (in L)}} \\ 22.99 & & V \text{ (in L)} = \frac{\# \text{ of mol}}{M} \\ 35.45 & & \\ \hline 58.44 & & \frac{0.089 \text{ mol}}{0.25 \text{ M}} = \boxed{0.36 \text{ L}} \end{aligned}$$

Part A Describing Molar Concentration

**Dick Bedeaux, Signe Kjelstrup, Jan V.
Sengers**



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Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

Biothermodynamics Part A Michael L. Johnson, Jo M. Holt, Gary K. Ackers, 2009-03-14 In the past several years there has been an explosion in the ability of biologists molecular biologists and biochemists to collect vast amounts of data on their systems This volume presents sophisticated methods for estimating the thermodynamic parameters of specific protein protein protein DNA and small molecule interactions The use of thermodynamics in biological research is used as an energy book keeping system While the structure and function of a molecule is important it is equally important to know what drives the energy force These methods look to answer What are the sources of energy that drive the function Which of the pathways are of biological significance As the base of macromolecular structures continues to expand through powerful techniques of molecular biology such as X ray crystal data and spectroscopy methods the importance of tested and reliable methods for answering these questions will continue to expand as well

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