

Spring Chapter 6

1. Consider the following energy level diagram for a reaction.

(Energy)

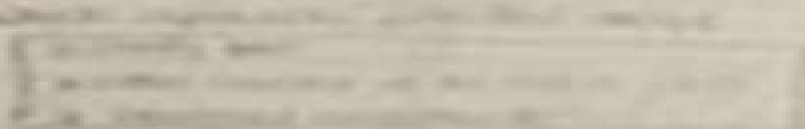
at reactants is 100 kJ

(kJ/mol)

at products is 200 kJ

(kJ/mol)

2. Which of the following is the correct energy diagram for the reaction?



3. A reaction has an activation energy of 100 kJ/mol. The reactants have an energy of 100 kJ/mol. What is the energy of the products?

a. 0 kJ/mol

b. 100 kJ/mol

c. 200 kJ/mol

4. Which of the following is the correct energy diagram for the reaction?

a. Reactants at 100 kJ/mol

b. Reactants at 200 kJ/mol

c. Reactants at 100 kJ/mol

d. Reactants at 200 kJ/mol

5. Consider the reaction of hydrogen and oxygen to form water. The reaction is exothermic. Which of the following is the correct energy diagram for the reaction?

a. Reactants at 100 kJ/mol

b. Reactants at 200 kJ/mol

c. Reactants at 100 kJ/mol

d. Reactants at 200 kJ/mol

6. 21.40 g of a sample of a substance is heated from 25.0°C to 100.0°C. The heat of fusion is 33.4 kJ/mol. What is the specific heat of the substance?

a. 0.900 J/g°C

b. 0.900 J/g°C

c. 0.900 J/g°C

d. 0.900 J/g°C

Sapling Learning Answers General Chemistry Chapters

RJ Shavelson



Sapling Learning Answers General Chemistry Chapters:

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