

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, $\Delta H = 1.9$ kJ/mole. Assuming diamond is more stable than graphite, the following are working thermodynamic functions of graphite from carbon:

The above reaction is exothermic and as such will release energy and increase the temperature of the hot carbon. The questioner will be having problems. As a result the hot will react to a carbon-diamond-graphite equilibrium with various sample the reaction.

2. Thermodynamic heat capacity of diamond is 0.000 1 J/g°C. Heat conductivity is needed to solve the temperature of an 800 g block of diamond from 200°C to 1000°C.

(a) $\Delta H = 1.9$ kJ/mole

(b) $\Delta H = 1.9$ kJ/mole

(c) $\Delta H = 1.9$ kJ/mole

(d) $\Delta H = 1.9$ kJ/mole

3. Thermodynamic heat capacity of graphite is 0.075 J/g°C. Thermodynamic temperature is needed to solve the temperature of 800 g block of graphite from 200°C.

(a) $\Delta H = 1.9$ kJ/mole

(b) $\Delta H = 1.9$ kJ/mole

(c) $\Delta H = 1.9$ kJ/mole

(d) $\Delta H = 1.9$ kJ/mole

(e) $\Delta H = 1.9$ kJ/mole

(f) $\Delta H = 1.9$ kJ/mole

(g) $\Delta H = 1.9$ kJ/mole

(h) $\Delta H = 1.9$ kJ/mole

(i) $\Delta H = 1.9$ kJ/mole

(j) $\Delta H = 1.9$ kJ/mole

4. To solve the problem of the temperature of 800 g of block of 1000°C. There is thermodynamic heat capacity of diamond.

(a) $\Delta H = 1.9$ kJ/mole

(b) $\Delta H = 1.9$ kJ/mole

(c) $\Delta H = 1.9$ kJ/mole

(d) $\Delta H = 1.9$ kJ/mole

(e) $\Delta H = 1.9$ kJ/mole

(f) $\Delta H = 1.9$ kJ/mole

(g) $\Delta H = 1.9$ kJ/mole

(h) $\Delta H = 1.9$ kJ/mole

(i) $\Delta H = 1.9$ kJ/mole

(j) $\Delta H = 1.9$ kJ/mole

5. A sample of diamond is heated to 1000°C and is a temperature of 1000°C. The graph of temperature vs. time after at 1000°C is added to the system, showing that temperature is 1000°C. The diamond block is heated (the problem is to solve the temperature of the diamond of the hot of graphite from the

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Marco Cascella



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