

Predicting Products of Chemical Reactions

This worksheet is designed to help you predict products of simple reactions of the four basic reaction types (synthesis, decomposition, single replacement, and double replacement) and combustion reactions.

For the first few reactions, the type of reaction is listed; you should predict the products, then balance. Further questions just have the reactants listed and you should decide on the type of reaction, as well as the correct products. Many of these reactions fall into the category of "redox" reactions, though do not let that confuse you...each can be described in terms of the four basic reaction types (except the combustion reactions).

Although states (s, l, g, aq) of the reactants and products are very important in a chemical reaction, don't worry about determining those for these problems. Rather, focus on what products might result from the reactants given. Pay particular attention to the ionic charge of species that you know form ions with only one possible charge (e.g., alkali metals, alkaline earth metals, halogens, etc.)

- a. Combustion: C_2H_6 + O_2 →
- b. Combustion: C_2H_4 + O_2 →
- c. Combustion: $\text{C}_2\text{H}_5\text{O}_2$ + O_2 →
1. Synthesis: Hg + S_8 →
2. Double displacement: CuCl_2 + H_2S →
3. Double displacement: NaOH + HClO_4 →
4. Decomposition: ZnCO_3 + heat →
5. Single replacement: HCl + Zn →
6. _____ Na + MgCl_2 →
7. _____ CaCl_2 + K_2CO_3 →
8. _____ K + Cl_2 →
9. _____ BaCl_2 + K_3PO_4 →
10. _____ H_2SO_4 + KOH →
11. _____ $\text{Al}_2(\text{CO}_3)_3$ + heat →
12. _____ Al + O_2 →
13. _____ $\text{Pt}(\text{NO}_3)_2$ + KOH →
14. _____ H_2SO_4 + BaCl_2 →
15. _____ Ca + AgCl →
16. _____ H_3PO_4 + FeBr_3 →
17. _____ Li + N_2 →
18. _____ HCl + $\text{Mg}(\text{OH})_2$ →
19. _____ $\text{Mg}(\text{OH})_2$ + heat →
20. _____ $\text{Fe}(\text{OH})_3$ + heat →

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Table of Contents Reaction Products Worksheet Answers Chemfiesta

1. Understanding the eBook Reaction Products Worksheet Answers Chemfiesta
 - The Rise of Digital Reading Reaction Products Worksheet Answers Chemfiesta
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Products Worksheet Answers Chemfiesta
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Products Worksheet Answers Chemfiesta
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Products Worksheet Answers Chemfiesta
 - Personalized Recommendations
 - Reaction Products Worksheet Answers Chemfiesta User Reviews and Ratings
 - Reaction Products Worksheet Answers Chemfiesta and Bestseller Lists
5. Accessing Reaction Products Worksheet Answers Chemfiesta Free and Paid eBooks

- Reaction Products Worksheet Answers Chemfiesta Public Domain eBooks
- Reaction Products Worksheet Answers Chemfiesta eBook Subscription Services
- Reaction Products Worksheet Answers Chemfiesta Budget-Friendly Options
- 6. Navigating Reaction Products Worksheet Answers Chemfiesta eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Products Worksheet Answers Chemfiesta Compatibility with Devices
 - Reaction Products Worksheet Answers Chemfiesta Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Products Worksheet Answers Chemfiesta
 - Highlighting and Note-Taking Reaction Products Worksheet Answers Chemfiesta
 - Interactive Elements Reaction Products Worksheet Answers Chemfiesta
- 8. Staying Engaged with Reaction Products Worksheet Answers Chemfiesta
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Products Worksheet Answers Chemfiesta
- 9. Balancing eBooks and Physical Books Reaction Products Worksheet Answers Chemfiesta
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Products Worksheet Answers Chemfiesta
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Products Worksheet Answers Chemfiesta
 - Setting Reading Goals Reaction Products Worksheet Answers Chemfiesta
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Products Worksheet Answers Chemfiesta
 - Fact-Checking eBook Content of Reaction Products Worksheet Answers Chemfiesta
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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