

Chapter 1: REACTION KINETICS

1.1 REACTION KINETICS

- a) Arrhenius equation: $k = A e^{-E_a/RT}$
- b) Transition state theory: $k = \frac{k_B T}{h} e^{-E_a/RT}$
- c) Reaction coordinate diagram: E_a is activation energy
- d) Rate law: $r = k [A]^m [B]^n$
- e) Order of reaction: $m+n$
- f) Half-life: $t_{1/2}$
- g) Integrated rate laws: $\ln[A] = \ln[A_0] - kt$
- h) Arrhenius plot: $\ln k = \ln A - E_a/RT$
- i) Reaction mechanism: $r = k [A]^m [B]^n$
- j) Rate of reaction: $r = \frac{1}{\nu} \frac{d[A]}{dt}$
- k) Rate of reaction: $r = \frac{1}{\nu} \frac{d[A]}{dt}$
- l) Rate of reaction: $r = \frac{1}{\nu} \frac{d[A]}{dt}$
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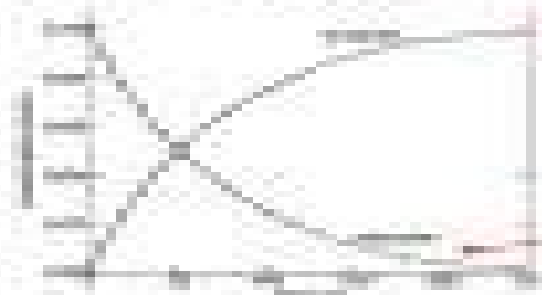
REACTION RATE

- a) Reaction rate: $r = \frac{1}{\nu} \frac{d[A]}{dt}$
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Reaction rate is the change in concentration of a reactant or product per unit time.

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Equilibrium is reached when the concentration of reactants and products are equal.

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Puja Mehta

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