



What you will do

Self-Test 2.1

Try to check how much you have learned in this lesson by answering the following questions.

I. Write the equilibrium expression for the following reactions

1. $\text{Ni}_{(s)} + 4\text{CO}_{(g)} \rightleftharpoons \text{Ni(CO)}_{4(g)}$
2. $\text{H}_2\text{O}_{(g)} \rightleftharpoons \text{H}_2\text{O}_{(l)}$
3. $\text{NiO}_{(s)} + \text{CO}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + \text{Ni}_{(s)}$
4. $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$
5. $3\text{H}_{2(g)} + \text{N}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$

II. Consider the exothermic reaction



Write $\rightarrow$ if the reaction will go forward, $\leftarrow$ for reverse and $\rightleftharpoons$ for no change in the direction of the reaction.

- | | |
|-------|-----------------------------|
| _____ | 1. Adding oxygen |
| _____ | 2. Removing sulfur dioxide |
| _____ | 3. Removing sulfur trioxide |
| _____ | 4. Adding helium gas |
| _____ | 5. Adding catalyst |



Key to answers on page 15.

Lesson 3. Applications of Chemical Equilibrium

In this lesson, we will study some of the wonderful applications of chemical equilibrium. The principles of chemical equilibrium are evident from the production of industrial chemicals, the acid-balance in lakes and streams, to the maintenance of our blood pH.

Haber Process: Ammonia synthesis

We have encountered the synthesis of ammonia from previous lessons. This time, we will take a closer look at this reaction. The complete equation is written as



Reaction Rates And Equilibrium Lesson Check Answers

JR Anderson



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Table of Contents Reaction Rates And Equilibrium Lesson Check Answers

1. Understanding the eBook Reaction Rates And Equilibrium Lesson Check Answers
 - The Rise of Digital Reading Reaction Rates And Equilibrium Lesson Check Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates And Equilibrium Lesson Check Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rates And Equilibrium Lesson Check Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates And Equilibrium Lesson Check Answers
 - Personalized Recommendations
 - Reaction Rates And Equilibrium Lesson Check Answers User Reviews and Ratings

- Reaction Rates And Equilibrium Lesson Check Answers and Bestseller Lists
- 5. Accessing Reaction Rates And Equilibrium Lesson Check Answers Free and Paid eBooks
 - Reaction Rates And Equilibrium Lesson Check Answers Public Domain eBooks
 - Reaction Rates And Equilibrium Lesson Check Answers eBook Subscription Services
 - Reaction Rates And Equilibrium Lesson Check Answers Budget-Friendly Options
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 - Reaction Rates And Equilibrium Lesson Check Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Reaction Rates And Equilibrium Lesson Check Answers
 - Interactive Elements Reaction Rates And Equilibrium Lesson Check Answers
- 8. Staying Engaged with Reaction Rates And Equilibrium Lesson Check Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rates And Equilibrium Lesson Check Answers
- 9. Balancing eBooks and Physical Books Reaction Rates And Equilibrium Lesson Check Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates And Equilibrium Lesson Check Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Rates And Equilibrium Lesson Check Answers
 - Setting Reading Goals Reaction Rates And Equilibrium Lesson Check Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Rates And Equilibrium Lesson Check Answers
 - Fact-Checking eBook Content of Reaction Rates And Equilibrium Lesson Check Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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