

Figure 1. Structure of Quinine.

# Experimental

# Reagents

- Sulfuric acid, H<sub>2</sub>SO<sub>4</sub>, high purity
- Potassium bromide, KBr, high purity
- Potassium iodide, KI, high purity
- Quinine, C<sub>20</sub>H<sub>24</sub>N<sub>2</sub>O<sub>2</sub>, high purity

#### Solutions

- sulfuric acid, 0.050 M. 3.0 mL of concentrated sulfuric acid was quantitatively transferred into a flask and diluted with 500 mL of distilled water. After vigorous stirring, the solution was transferred into a 1 L flask and diluted with distilled water to the mark.
- stock quinine,  $1.25 \times 10^3$
- µg/mL. 0.3125 g of quinine were dissolved in 100 mL of 0.050 M sulfuric acid solution. After vigorous stirring, an additional amount of sulfuric acid was added until the mark was reached in 250 mL.
- working standard quinine. 1.0 ×10<sup>-2</sup>
- µg/mL. 2 mL of stock quinine were quantitatively transferred to a 250 mL flask and brought to the mark with a 0.050 M sulfuric acid solution.
- potassium bromide,  $2.5 \times 10^{-2}$

# **Quinine In Tonic Water Report**

**Chao Zhang** 

**Quinine In Tonic Water Report:** 

The Enigmatic Realm of **Quinine In Tonic Water Report**: Unleashing the Language is Inner Magic

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