

En un haluro de alquilo el átomo de halógeno está enlazado a un átomo de carbono con hibridación  $sp^3$ . El halógeno es más electronegativo que el carbono, y el enlace  $C-X$  está polarizado con una carga parcial positiva en el carbono y una carga parcial negativa en el halógeno.

## 6.4 Estructura de los haluros de alquilo



$$\mu = 4.8 \times \delta \times d$$

donde  $\delta$  es la carga y  $d$  la longitud de enlace

Las electronegatividades de los halógenos *aumentan* en el siguiente orden:

|                    |     |   |     |   |     |   |     |
|--------------------|-----|---|-----|---|-----|---|-----|
|                    | I   | < | Br  | < | Cl  | < | F   |
| electronegatividad | 2.7 |   | 3.0 |   | 3.2 |   | 4.0 |

Las longitudes de enlace carbono-halógeno *aumentan* a medida que aumenta el volumen de los átomos de halógeno (mayor radio atómico) en el orden:

|                     |        |   |        |   |        |   |        |
|---------------------|--------|---|--------|---|--------|---|--------|
|                     | C—F    | < | C—Cl   | < | C—Br   | < | C—I    |
| longitud de enlace: | 1.38 Å |   | 1.78 Å |   | 1.94 Å |   | 2.14 Å |

Estos dos efectos son opuestos, ya que los halógenos que tienen más volumen y mayor longitud de enlace tienen las electronegatividades más pequeñas. El resultado total es que los momentos dipolares de enlace *aumentan* en el orden:

|                    |        |   |        |   |        |   |        |
|--------------------|--------|---|--------|---|--------|---|--------|
|                    | C—I    | < | C—Br   | < | C—F    | < | C—Cl   |
| momento dipolar, D | 1.29 D |   | 1.48 D |   | 1.51 D |   | 1.56 D |

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## **Table of Contents Quimicaorganica Webcindario Com**

1. Understanding the eBook Quimicaorganica Webcindario Com
  - The Rise of Digital Reading Quimicaorganica Webcindario Com
  - Advantages of eBooks Over Traditional Books
2. Identifying Quimicaorganica Webcindario Com
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Quimicaorganica Webcindario Com
  - User-Friendly Interface
4. Exploring eBook Recommendations from Quimicaorganica Webcindario Com
  - Personalized Recommendations
  - Quimicaorganica Webcindario Com User Reviews and Ratings
  - Quimicaorganica Webcindario Com and Bestseller Lists
5. Accessing Quimicaorganica Webcindario Com Free and Paid eBooks
  - Quimicaorganica Webcindario Com Public Domain eBooks
  - Quimicaorganica Webcindario Com eBook Subscription Services
  - Quimicaorganica Webcindario Com Budget-Friendly Options

6. Navigating Quimicaorganica Webcindario Com eBook Formats
  - ePub, PDF, MOBI, and More
  - Quimicaorganica Webcindario Com Compatibility with Devices
  - Quimicaorganica Webcindario Com Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Quimicaorganica Webcindario Com
  - Highlighting and Note-Taking Quimicaorganica Webcindario Com
  - Interactive Elements Quimicaorganica Webcindario Com
8. Staying Engaged with Quimicaorganica Webcindario Com
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Quimicaorganica Webcindario Com
9. Balancing eBooks and Physical Books Quimicaorganica Webcindario Com
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Quimicaorganica Webcindario Com
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Quimicaorganica Webcindario Com
  - Setting Reading Goals Quimicaorganica Webcindario Com
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Quimicaorganica Webcindario Com
  - Fact-Checking eBook Content of Quimicaorganica Webcindario Com
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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