

Pre-Lab Information

Purpose	Explore the rate of a chemical reaction at different temperatures and with different reactant particle sizes.
Time	Approximately 45 minutes
Question	What are the effects of temperature and a reactant's particle size on reaction rate?
Hypothesis 1	If you increase the temperature of a reaction, then the reaction rate will increase because particles experience more collisions at higher temperatures.
Hypothesis 2	If you decrease the particle size of a reactant, then the reaction rate will increase because more of the reactants' surface area is exposed allowing more particles to make contact with each other.
Variables	Independent variables: particle size, temperature Dependent variable: reaction rate
Summary	The reaction between effervescent tablets and water will produce CO_2 gas, which you can observe in the form of bubbles; when the bubbling stops, the reaction is complete. You will measure the reaction times and rates for tablets dissolved at four temperatures and broken into three different particle sizes.

Safety

- Always wear a lab coat and safety goggles when performing an experiment.
- Behavior in the lab needs to be purposeful. Use caution when using the hot plate, as it can cause burns.
- Check glassware such as graduated cylinders and beakers for cracks and chips prior to use.
- Use the right gear, such as chemically resistant gloves and oven mitts, when performing the experiment.
- Report all accidents—no matter how big or small—to your teacher.

Procedure

Step 1: Gather Materials

- 250 mL graduated cylinder
- Thermometer
- Water
- Timer
- Four 250 mL beakers
- Seven 1,000 mg effervescent tablets
- Two pieces of filter paper
- 600 mL beaker
- Ice
- Hot plate

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TD Snyder



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