

Calculating Average Rates of Reaction

You are required to show work for all calculations.

1. During the combustion of methane, CH_4 , shown by the reaction



The concentration of methane was measured at various time intervals and the following results were obtained:

Time (s)	$[\text{CH}_4]$ ($\text{mol} \cdot \text{l}^{-1}$)
10	2.40
20	1.20
30	0.80
40	0.60

Calculate the average rate of loss of methane during the 10 to 40 second time period.

Solution:

$$\begin{aligned}
 \text{rate} &= \frac{\Delta[\text{CH}_4]}{\Delta t} \\
 &= \frac{[0.60 - 2.4]\text{mol}}{(40 - 10)\text{s} \times \text{L}} \\
 &= \frac{-1.8\text{mol}}{30\text{s} \times \text{L}} \\
 &= -0.06 \frac{\text{mol}}{\text{L} \times \text{s}}
 \end{aligned}$$

Note: the answer is negative because we are calculating the **rate of loss of methane**.

Calculating Reaction Rates using Stoichiometry

Example:

In the following decomposition reaction,



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In some sort of inundated with screens and the cacophony of instantaneous communication, the profound energy and mental resonance of verbal beauty frequently fade into obscurity, eclipsed by the regular barrage of sound and distractions. Yet, nestled within the musical pages of **Reaction Rate Practice Problems**, a fascinating perform of fictional splendor that impulses with fresh thoughts, lies an unforgettable trip waiting to be embarked upon. Written with a virtuoso wordsmith, this magical opus courses readers on a mental odyssey, lightly exposing the latent possible and profound impact stuck within the complex web of language. Within the heart-wrenching expanse with this evocative evaluation, we will embark upon an introspective exploration of the book is central subjects, dissect their captivating publishing fashion, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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