

CHAPTER 13

Rates of Reaction

■ SOLUTIONS TO EXERCISES

Acetaminophen is a painkiller that gives analgesic results in the treatment of headaches, whereas aspirin can relieve pain, but also irritate the stomach. The drug content of each tablet is the amount measured gravimetrically. In determining the actual amount of each drug present each manufacturer reports a weight, not the actual amount but the theoretical.

- 13.1. Data is presented in Table 13.1. Data values based on $\text{g} \cdot \text{mol}^{-1}$ values are shown with units for convenience of determining a useful equation.

$$\text{Aspirin} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass}$$

- 13.2. Data is presented in Table 13.2. Data values are presented in $\text{g} \cdot \text{mol}^{-1}$ and $\text{g} \cdot \text{mol}^{-1}$.

- 13.3. The data table compares the actual and theoretical values of the drug content of each tablet. The data values are presented in Table 13.3.

- 13.4. The data table compares the actual and theoretical values of the drug content of each tablet. The data values are presented in Table 13.4.

Data is presented in

The data table compares the actual and theoretical values of the drug content of each tablet. The data values are presented in Table 13.5.

$$\text{Aspirin} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass}$$

- 13.5. Data is presented in Table 13.6. Data values are presented in $\text{g} \cdot \text{mol}^{-1}$ and $\text{g} \cdot \text{mol}^{-1}$.

$$\text{Aspirin} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass}$$

Data is presented in Table 13.7.

$$\text{Aspirin} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass}$$

Data is presented in

$$\text{Aspirin} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass} = \frac{\text{Weight}}{\text{Molar Mass}} \times \text{Molar Mass}$$

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