CHAPTER 16 REVIEW

Reaction Energy

SECTION 2

SHORT ANSWER Answer the following questions in the space provided.

 For the following examples, state whether the change in entropy favors the forward or reverse reaction:

forward reaction a. $HCl(I) \rightleftharpoons HCl(g)$

reverse reaction b. $C_6H_{12}O_6(aq) \rightleftharpoons C_6H_{12}O_6(s)$

forward reaction c. $2NH_3(g) \rightleftharpoons N_2(g) + 3H_2(g)$

reverse reaction d. $3C_2H_4(g) \rightleftharpoons C_6H_{12}(I)$

2. (change in free energy)G = (change in catropy)S
Write an equation that shows the relationship between enthalpy.
ΔH, entropy, ΔS, and free energy, ΔG.

b. For a reaction to occur spontaneously, the sign of ΔG should be _____

a. The enthalpy factor favors the forward reaction. True or False?

b. The sign of $T\Delta S^o$ is negative. This means the entropy factor favors the

c. Given that ΔG° for the above reaction in the forward direction is positive, which term is greater in magnitude and therefore predominates, TΔS or ΔH?

T(change in)S is greater than (change in)H; (change in)H is negative, but [F]G is positive, indicating the reaction is not spontaneous. The randomness of the reactants and the temperature of the reaction overshadow the exothermic factor.

4. Consider the following equation for the vaporization of water:

 $H_2O(I) \rightleftharpoons H_2O(g)$ $\Delta H = +40.65 \text{ kJ/mol at } 100^{\circ}\text{C}$

endothermic a. Is the forward reaction exothermic or endothermic?

b. Does the enthalpy factor favor the forward or reverse reaction?

forward reaction

c. Does the entropy factor favor the forward or reverse reaction?

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