

HALF YEARLY EXAMINATION — 2014-15

CHEMISTRY

Class – XI

Time Allowed : 3 Hrs.

Maximum Marks : 70

General Instructions :

- (i) Attempt all questions.
- (ii) Q 1-5 are very short answer type questions carrying 1 mark each.
- (iii) Q 6-10 are short answer type questions carrying 2 marks each.
- (iv) Q 11-22 are also short answer type questions carrying 3 marks each.
- (v) Q 23 is a value based question carrying 4 marks.
- (vi) Q 24-26 are long answer type questions carrying 5 marks each.
- (vii) There is no overall choice in the paper. However, there is an internal choice in some questions.
- (viii) Use of calculator is prohibited. Log tables can be used if necessary.

1. Give an example of a molecule which has intramolecular hydrogen bonding.
2. Write the electronic configuration of La ($z = 57$).
3. Which of the following species has the largest and smallest size? Mg, Mg^{2+} , Al, Al^{3+}
4. What is the significance of van der Waals parameters 'a' and 'b'?
5. What will happen to the wavelength associated with the moving particle if its velocity is doubled?
6. Two bulbs B_1 and B_2 of equal capacity contains 10g oxygen (O_2) and ozone (O_3) respectively. Which bulb will have greater number of oxygen atoms?
(Atomic mass of oxygen = 16 u and Avogadro's number = 6.022×10^{23})
7. Using ideal gas equation, show that at a given temperature, density of a gas is directly proportional to the pressure of the gas.
8. Derive a relation between C_p and C_v for one mole of an ideal gas.
9. Explain dispersion forces with the help of a suitable example.

OR

The Maxwell-Boltzmann distribution curve for a gas at three different temperatures is shown below. Answer the questions based on it.

- (a) The value of the most probable speed (c^*) obtained at temperature T_3 is greater than that obtained at temperature T_1 and T_2 . Justify.

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M Mark



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Table of Contents Question Papers Chemistry N4 July 2014

1. Understanding the eBook Question Papers Chemistry N4 July 2014
 - The Rise of Digital Reading Question Papers Chemistry N4 July 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Question Papers Chemistry N4 July 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Question Papers Chemistry N4 July 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Question Papers Chemistry N4 July 2014
 - Personalized Recommendations

- Question Papers Chemistry N4 July 2014 User Reviews and Ratings
- Question Papers Chemistry N4 July 2014 and Bestseller Lists
- 5. Accessing Question Papers Chemistry N4 July 2014 Free and Paid eBooks
 - Question Papers Chemistry N4 July 2014 Public Domain eBooks
 - Question Papers Chemistry N4 July 2014 eBook Subscription Services
 - Question Papers Chemistry N4 July 2014 Budget-Friendly Options
- 6. Navigating Question Papers Chemistry N4 July 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Question Papers Chemistry N4 July 2014 Compatibility with Devices
 - Question Papers Chemistry N4 July 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Question Papers Chemistry N4 July 2014
 - Highlighting and Note-Taking Question Papers Chemistry N4 July 2014
 - Interactive Elements Question Papers Chemistry N4 July 2014
- 8. Staying Engaged with Question Papers Chemistry N4 July 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Question Papers Chemistry N4 July 2014
- 9. Balancing eBooks and Physical Books Question Papers Chemistry N4 July 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Question Papers Chemistry N4 July 2014
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Question Papers Chemistry N4 July 2014
 - Setting Reading Goals Question Papers Chemistry N4 July 2014
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Question Papers Chemistry N4 July 2014
 - Fact-Checking eBook Content of Question Papers Chemistry N4 July 2014

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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