

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES  
(CHEMISTRY)**

**TABLE 1      PHYSICAL CONSTANTS**

NAME	SYMBOL	VALUE
Magnitude of charge on electron	$e$	$1,6 \times 10^{-19} \text{ C}$
Mass of an electron	$m_e$	$9,1 \times 10^{-31} \text{ kg}$
Standard pressure	$p^\circ$	$1,01 \times 10^5 \text{ Pa}$
Molar gas volume at STP	$V_m$	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature	$T^\circ$	$273 \text{ K}$
Avogadro's constant	$N_A$	$6,02 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	$F$	$96\,500 \text{ C} \cdot \text{mol}^{-1}$

**TABLE 2      CHEMISTRY FORMULAE**

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$	$n = \frac{V}{V_m}$
$c = \frac{n}{V}$ OR $c = \frac{m}{MV}$	$K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$ at 25 °C (298 K)	
$q = It$ $q = nF$	$E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$ $E_{\text{cell}}^\circ = E_{\text{oxidising agent}}^\circ - E_{\text{reducing agent}}^\circ$	

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