## QUESTION 10 (Start on a new page.)

The balanced chemical equation for the EXTRACTION of gold from its ore is as follows:

$$4Au(s) + 8NaCN(aq) + 2H2O(l) + O2(g) \rightarrow 4NaAu(CN)2(aq) + 4NaOH(aq)$$

- 10.1 State ONE disadvantage of using cyanide (CN') in the extraction of gold. (1)
- 10.2 Will the final solution of the extraction process be ACIDIC or BASIC (ALKALINE)? Give a reason for the answer.
  (2)
- 10.3 Determine the oxidation number of gold in NaAu(CN)<sub>2</sub>. (1)
- 10.4 Write down the FORMULA of the reducing agent in the reaction above. (1)

Zinc powder is now used to PRECIPITATE the gold.

The balanced equation for the reaction is:

$$Zn(s) + 2NaAu(CN)_2(aq) \rightarrow 2Au(s) + Zn(CN)_2(aq) + 2NaCN(aq)$$

- 10.5 Does zinc undergo OXIDATION or REDUCTION during the precipitation reaction? (1)
- 10.6 Write down a half-reaction to support the answer to QUESTION 10.5. (2)
- 10.7 Calculate the percentage of gold in NaAu(CN)<sub>2</sub>. (2)
  [10]

TOTAL: 150

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