



NATIONAL  
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2  
(CHEMISTRY)  
EXEMPLAR

MARKS: 100  
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.  
Use blue or black ink for writing. Do not use red ink.

## QUESTION 3

# PHYSICAL SCIENCES PAPER 2 INTERMOLECULAR FORCES

**GRADE 11**

**NOVEMBER 2020**

Examiner's use: \_\_\_\_\_ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the  $\Delta H_{\text{vap}}$  and  $\Delta H_{\text{fusion}}$  for 1-hexanol in kJ mol<sup>-1</sup> and 1 mol of 1-hexanol.

Compound	Molecular mass (M <sub>r</sub> )	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B or of A, B and C or of others as well? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and is the prime reason of their volatility. (2)
- 3.1.4 What factor of the compounds A and B is responsible for the difference in their boiling points? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Which compound (A or B) will have the higher surface tension? (2)

Compound C (M<sub>r</sub> 102) will also be in the table. Use a suitable intermolecular force that both compounds A and B do not have as a secondary factor during point of 3.1.2.

- 3.1.8 Explain why the boiling point of compound B is HIGHER than that of compound A and B by referring to the  $\Delta H_{\text{vap}}$  and intermolecular or intramolecular forces involved. (2)

3.2 Add column C to the table in your answer of question 3.1.2.  $\Delta H_{\text{fusion}}$  and  $\Delta H_{\text{vap}}$  are given in kJ mol<sup>-1</sup> and 1 mol of the respective compounds.

Compound	Molecular mass (M <sub>r</sub> )	Boiling point (°C)	$\Delta H_{\text{vap}}$ (kJ mol <sup>-1</sup> )	$\Delta H_{\text{fusion}}$ (kJ mol <sup>-1</sup> )
A	102	16,2	21,2	5,1
B	102	21	21,2	5,1

- 3.2.1 Is solid pure 1-hexanol at 0°C, all the solid dissolved? (2)
- 3.2.2 Explain the process in QUESTION 3.2.1 about the melting in the intermolecular forces involved. (2)

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# Physical Science Paper 2 Intermolecular Forces Investigation

**Lei Huang**



## **Physical Science Paper 2 Intermolecular Forces Investigation:**

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