

**Balancing Equations Answers**

1.  $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
2.  $\text{H}_3\text{PO}_4 + 3 \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
3.  $6 \text{K} + \text{B}_2\text{O}_3 \rightarrow 3 \text{K}_2\text{O} + 2 \text{B}$
4.  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
5.  $10 \text{Na} + 2 \text{NaNO}_3 \rightarrow 6 \text{Na}_2\text{O} + \text{N}_2$
6.  $4 \text{C} + \text{S}_8 \rightarrow 4 \text{CS}_2$
7.  $2 \text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}_2$
8.  $2 \text{N}_2 + 5 \text{O}_2 \rightarrow 2 \text{N}_2\text{O}_5$
9.  $2 \text{H}_3\text{PO}_4 + 3 \text{Mg(OH)}_2 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$
10.  $2 \text{NaOH} + \text{H}_2\text{CO}_3 \rightarrow \text{Na}_2\text{CO}_3 + 2 \text{H}_2\text{O}$
11.  $\text{KOH} + \text{HBr} \rightarrow \text{KBr} + \text{H}_2\text{O}$
12.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
13.  $4 \text{Na} + \text{O}_2 \rightarrow 2 \text{Na}_2\text{O}$
14.  $2 \text{Al(OH)}_3 + 3 \text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + 6 \text{H}_2\text{O}$
15.  $16 \text{Al} + 3 \text{S}_8 \rightarrow 8 \text{Al}_2\text{S}_3$
16.  $6 \text{Cs} + \text{N}_2 \rightarrow 2 \text{Cs}_3\text{N}$
17.  $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
18.  $10 \text{Rb} + 2 \text{RbNO}_3 \rightarrow 6 \text{Rb}_2\text{O} + \text{N}_2$
19.  $2 \text{C}_6\text{H}_6 + 15 \text{O}_2 \rightarrow 12 \text{CO}_2 + 6 \text{H}_2\text{O}$
20.  $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$
21.  $2 \text{C}_{10}\text{H}_{22} + 31 \text{O}_2 \rightarrow 20 \text{CO}_2 + 22 \text{H}_2\text{O}$
22.  $\text{Al(OH)}_3 + 3 \text{HBr} \rightarrow \text{AlBr}_3 + 3 \text{H}_2\text{O}$
23.  $2 \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
24.  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
25.  $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
26.  $3 \text{Li} + \text{AlCl}_3 \rightarrow 3 \text{LiCl} + \text{Al}$
27.  $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
28.  $3 \text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + 3 \text{H}_2\text{O}$
29.  $3 \text{Rb} + \text{P} \rightarrow \text{Rb}_3\text{P}$
30.  $\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$
31.  $2 \text{Al(OH)}_3 + 3 \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 6 \text{H}_2\text{O}$
32.  $2 \text{Na} + \text{Cl}_2 \rightarrow 2 \text{NaCl}$
33.  $16 \text{Rb} + \text{S}_8 \rightarrow 8 \text{Rb}_2\text{S}$
34.  $2 \text{H}_3\text{PO}_4 + 3 \text{Ca(OH)}_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$
35.  $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
36.  $2 \text{Li} + 2 \text{H}_2\text{O} \rightarrow 2 \text{LiOH} + \text{H}_2$
37.  $\text{Ca}_3(\text{PO}_4)_2 + 3 \text{SiO}_2 + 5 \text{C} \rightarrow 3 \text{CaSiO}_3 + 5 \text{CO} + 2 \text{P}$
38.  $4 \text{NH}_3 + 3 \text{O}_2 \rightarrow 2 \text{N}_2 + 6 \text{H}_2\text{O}$
39.  $4 \text{FeS}_2 + 11 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
40.  $5 \text{C} + 2 \text{SO}_2 \rightarrow \text{CS}_2 + 4 \text{CO}$
41.  $\text{Fe} + \text{S} \rightarrow \text{FeS}$
42.  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
43.  $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
44.  $\text{Fe}_2\text{O}_3 + 3 \text{C} \rightarrow 3 \text{CO} + 2 \text{Fe}$
45.  $\text{K}_2\text{O} + \text{H}_2\text{O} \rightarrow 2 \text{KOH}$
46.  $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow 2 \text{KCl} + \text{BaCO}_3$
47.  $\text{Mg(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2 \text{H}_2\text{O}$
48.  $2 \text{KF} + \text{BaBr}_2 \rightarrow \text{BaF}_2 + 2 \text{KBr}$
49.  $\text{HCl} + \text{NH}_3 \rightarrow \text{NH}_4\text{Cl}$
50.  $\text{Bi}_2(\text{SO}_4)_3 + 6 \text{NH}_4\text{OH} \rightarrow 2 \text{Bi(OH)}_3 + 3 (\text{NH}_4)_2\text{SO}_4$

# Physical Science Balancing 1 Answer Key

**TD Snyder**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

## **Physical Science Balancing 1 Answer Key:**

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