#### **Chapter 4 - Practice Problems Solutions**

 Complete and balance the following reactions. Remember to form the stable product and then balance. Assume the product is solid. What type of reactions are these?

\_\_\_\_ Mg (s) + \_\_\_\_ O<sub>2</sub> (g) 
$$\rightarrow$$
 \_\_\_\_ Na (s) + \_\_\_\_ Cl<sub>2</sub> (g)  $\rightarrow$  2 Mg (s) + O<sub>2</sub> (g)  $\rightarrow$  2 Na (s) + Cl<sub>2</sub> (g)  $\rightarrow$  2 Na Cl (s)

\_\_\_\_ Al (s) + \_\_\_\_ O<sub>2</sub> (g) 
$$\rightarrow$$
 \_\_\_\_ Ca (s) + \_\_\_\_ Br<sub>2</sub> (g)  $\rightarrow$  4 Al (s) + 3 O<sub>2</sub> (g)  $\rightarrow$  2 Al<sub>2</sub>O<sub>3</sub> (s) Ca (s) + Br<sub>2</sub> (g)  $\rightarrow$  CaBr<sub>2</sub> (s)

These are all combination reactions.

2) Balance the following reactions. What type of reactions are these?

$$C_3H_8(g) + ___O_2(g) \Rightarrow __CO_2(g) + ___H_2O(g)$$
  
 $C_3H_8(g) + 5 O_2(g) \Rightarrow __CO_2(g) + 4 H_2O(g)$   
 $C_4H_8(g) + __O_2(g) \Rightarrow __CO_2(g) + __H_2O(g)$   
 $C_4H_8(g) + 15 O_2(g) \Rightarrow 12 CO_2(g) + 6 H_2O(g)$ 

These are both combustion reactions.

3) Balance the following reactions. What type of reactions are these?

\_\_\_\_\_ 
$$H_2O(I)$$
 → \_\_\_\_\_  $H_2(g)$  + \_\_\_\_\_  $O_2(g)$  \_\_\_\_\_  $KCIO_3(s)$  → \_\_\_\_\_  $KCI(s)$  + \_\_\_\_\_  $O_2(g)$   
2  $H_2O(I)$  → 2  $H_2(g)$  +  $O_2(g)$  2  $KCIO_3(s)$  → 2  $KCI(s)$  + 3  $O_2(g)$ 

These are both decomposition reactions.

4) Balance the following reactions. What type of reactions are these?

\_\_\_\_ Mg (s) + \_\_\_\_ HCl (aq) 
$$\rightarrow$$
 \_\_\_\_ MgCl<sub>2</sub> (aq) + \_\_\_\_ H<sub>2</sub> (g)  
Mg (s) + 2 HCl (aq)  $\rightarrow$  MgCl<sub>2</sub> (aq) + H<sub>2</sub> (g)  
\_\_\_\_ Ag (s) + \_\_\_\_ H<sub>2</sub>S (aq)  $\rightarrow$  \_\_\_\_ Ag<sub>2</sub>S (s) + \_\_\_\_ H<sub>2</sub> (g)  
2 Ag (s) + H<sub>3</sub>S (aq)  $\rightarrow$  Ag<sub>2</sub>S (s) + H<sub>2</sub> (g)

These are both single displacement reactions.

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**Mike Jess** 

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