Stoichiometry Problem Set A

Directions: Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

Part 1: Straight Stoichiometry Problems

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

For the reaction HNO_{Neq} + Mg(OH)₁₀₀ → Mg(NO₂)_{3(eq)} + H₂O₍₀₎, how many grams of magnesium nitrate are produced from 8.00 mol of nitric acid (HNO₂)? A

MM=148.39

For the reaction 2 KClO₃₀₀ → 2 KCl₀₀ + 3 O₂₀₀, how many moles of potassium chlorate are required to produce 160. g of oxygen? A

MH = 32.09

- KCID3
- 3.) For the reaction $2H_{200} + O_{200} \rightarrow 2H_2O_{50}$, how many grams of water are produces from 5.22 liters of oxygen gas?

MH= 18.09

In the reaction Na₍₀₎ + H₂O₍₀₎ → NaOH₍₀₋₀₎ + H₂₍₀₎, how many grams of hydrogen are produced if 1.33 grams of sodium react completely?

MH= 23.09

5.) In the reaction ____CO_{2(a)} + ___H₂O₀₀ → ___H₂CO_{2(a)}, how many moles of carbonic acid can be produced from 97.2 g of CO₂?

MH 44.09

2.21 moles Heca

Did you achieve our target goal?

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

Feeling VERY confused

3 I think I am catching on

Doing great! Bring it on!

Sample Problem Set Stoichometry Chemistry

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