

Focus 3: States of Matter

3A: The Nature of Gases

11. Pressure is the amount of force exerted by a gas over a certain area.
- $P = F/A$
 - The SI unit of pressure is the pascal, where $1 \text{ Pa} = 1 \text{ kg/m} \cdot \text{s}^2$
 - $1 \text{ bar} = 1.013 \times 10^5 \text{ Pa}$
 - $1 \text{ atm} = 1.01325 \times 10^5 \text{ Pa}$
 - $1 \text{ atm} = 760 \text{ Torr}$

3B: The Gas Laws

8. Boyle's Law states that gases placed at a constant temperature
- A. An Isotherm is a curve for a constant-temperature plot.
9. Boyle's Law pressure is inversely proportional to volume
- A. $P_1 V_1 = P_2 V_2$
10. Avogadro's Principle states that gas molecules occupy the same volume regardless of chemical identity, given that it is under constant pressure and temperature
11. The Ideal Gas Law summarizes the relation between pressure, volume, and amount of a substance
- A. $PN = nRT$
1. P = pressure
2. V = volume
3. n = moles
4. N = total gas constant which is equal to:

TABLE 35.3 The Gas Constant, R

$\mu_{\text{H}^+} = 1.0 \times 10^{-10} \text{ E}_{\text{vacuum}} \text{ (K}^{-1} \text{ cmol}^{-1})$
 $\mu_{\text{H}_2\text{O}} = 46.0 \times 10^{-10} \text{ E}_{\text{vacuum}} \text{ (K}^{-1} \text{ cmol}^{-1})$
 $\mu_{\text{H}_2\text{O}} = 45.2 \times 10^{-10} \text{ E}_{\text{vacuum}} \text{ (K}^{-1} \text{ cmol}^{-1})$
 $\mu_{\text{H}_2\text{O}} = 46.0 \times 10^{-10} \text{ E}_{\text{vacuum}} \text{ (K}^{-1} \text{ cmol}^{-1})$
 $\mu_{\text{H}_2\text{O}} = 46.0 \times 10^{-10} \text{ E}_{\text{vacuum}} \text{ (K}^{-1} \text{ cmol}^{-1})$

$$\frac{1}{2} \times \frac{1}{2} = \frac{1}{4}$$

- STP = standard temperature and pressure (0 degrees C and 1 atm)

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