

First Year Higher Secondary Examination

PART – III

CHEMISTRY

Maximum: 60 Scores

Time: 2 Hours

Cool-off time: 15 Minutes

1. (a) Determine the number of moles present in 0.55 mg of electrons.
 (i) 1 mole (ii) 2 mole (iii) 1.5 mole (iv) 0.5 mole (1)
 (b) Give the empirical formula of the following. (2)
 $C_8H_{12}O_6$, C_6H_6 , CH_3COOH , $C_6H_6Cl_6$
- (c) Two elements, carbon and hydrogen combine to form C_2H_6 , C_2H_4 and C_2H_2 , Identify the law illustrated here. (1)
2. (a) (i) Write the electronic configuration of chromium ($Z = 24$). (1)
 (ii) Find the number of electrons in the subshells with azimuthal quantum number $l = 2$. (1)
 (iii) Represent the orbital with quantum numbers $n = 1$ and $l = 0$ (1)
 (b) Give the mathematical representation of Heisenberg's uncertainty principle and its one important significance. (2)
3. Electron gain enthalpy is one of the important periodic property.
 (a) Define electron gain enthalpy. (1)
 (b) Explain any two factors affecting electron gain enthalpy. (2)
 (c) Write the oxidation state and co-valency of Al in $(AlF_6)^{3-}$. (1)
4. The geometry of the molecule is decided by type of hybridization.
 (a) Discuss the shape of PCl_5 molecules using hybridization. (2)
 (b) Give the reason for the high reactivity of PCl_5 . (2)
 (c) Isoelectric species have the same bond order. Among the following, choose the pair having same bond order.
 CN^- , $\bar{O}_2$, NO^+ , CN^+ . (1)

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