

Neutralization-Titration II

A. Molarity of an Unknown Acid

Example

Data Table

ONLY

	Sample 1		Sample 2		Sample 3 (if needed)	
	Acid*	Base	Acid*	Base	Acid*	Base
Final buret reading	X	20.0	18.0	37		
Initial buret reading	X	12.5	14.0	30		
Volume used	10 ml	7.5 ml		7.0 ml		

*If a pipet is used to measure the volume of acid, record only in the space for volume used.

Molarity of base (NaOH) 0.2 M

CALCULATIONS: In the spaces below, show calculation setups for Sample 1 only. Show answers for both samples in the boxes.

$$(V)(M) = (V) \left(\frac{\text{moles}}{L} \right) = \text{moles}$$

1. Moles of base (NaOH) (if needed)

$$(7.0 \text{ mL}) \rightarrow \left(\frac{1L}{1000 \text{ mL}} \right) (0.2 \frac{\text{moles}}{L}) =$$

2. Moles of acid used to neutralize (react with) the above number of moles of base

moles HCl = moles NaOH when HCl solution is neutralized

3. Molarity of acid

$$M_{\text{HCl}} = \frac{\text{moles HCl}}{L_{\text{HCl}}} =$$

4. Average molarity of acid

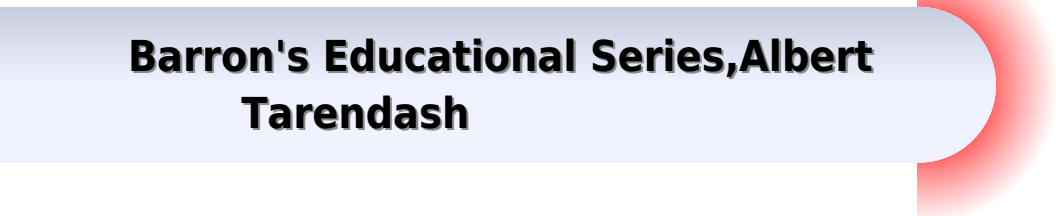
~~5. Unknown acid number~~

Sample 1	Sample 2	Sample 3 (if needed)
<u>0.00198</u>	<u>0.00198</u>	
<u>0.00198 M</u>	<u>0.002 M</u>	
<u>0.1994 M</u>	<u>0.196 M</u>	

0.1972

Neutralization And Titration Answers

**Barron's Educational Series, Albert
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Table of Contents Neutralization And Titration Answers

1. Understanding the eBook Neutralization And Titration Answers
 - The Rise of Digital Reading Neutralization And Titration Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Neutralization And Titration Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Neutralization And Titration Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Neutralization And Titration Answers
 - Personalized Recommendations
 - Neutralization And Titration Answers User Reviews and Ratings

- Neutralization And Titration Answers and Bestseller Lists
- 5. Accessing Neutralization And Titration Answers Free and Paid eBooks
 - Neutralization And Titration Answers Public Domain eBooks
 - Neutralization And Titration Answers eBook Subscription Services
 - Neutralization And Titration Answers Budget-Friendly Options
- 6. Navigating Neutralization And Titration Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Neutralization And Titration Answers Compatibility with Devices
 - Neutralization And Titration Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Neutralization And Titration Answers
 - Highlighting and Note-Taking Neutralization And Titration Answers
 - Interactive Elements Neutralization And Titration Answers
- 8. Staying Engaged with Neutralization And Titration Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Neutralization And Titration Answers
- 9. Balancing eBooks and Physical Books Neutralization And Titration Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Neutralization And Titration Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Neutralization And Titration Answers
 - Setting Reading Goals Neutralization And Titration Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Neutralization And Titration Answers
 - Fact-Checking eBook Content of Neutralization And Titration Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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