

EXPERIMENT 10: REPERCENT

GENERAL SINGLE-REPLACEMENT REACTIONS

Part A. Reaction of an active metal with water in a beaker.

1. Describe the metal, sodium. *Very active, highly soluble in water.*
2. Describe what happened when sodium was added to the water. *It reacted strongly. The solution became cloudy and the metal disappeared.*
3. Write a balanced equation for the reaction between sodium and water.
$$2\text{Na} + \text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2 \uparrow$$
4. Describe what happened when a burning splint was brought to the mouth of the test tube which contained the water. *Popping noise.*
5. Write an equation for the reaction that happened when the burning splint was brought to the mouth of the test tube.
$$\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$$

Part B. Describe what happened in each test if a reaction happened and write an equation for each. When "No Reaction" is indicated, write "No Reaction".

TEST 1	Aluminum	$2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2 \uparrow$
TEST 2	Zinc	$\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2 \uparrow$
TEST 3	Copper	$\text{Cu} + \text{H}_2\text{SO}_4 \rightarrow \text{CuSO}_4 + \text{H}_2 \uparrow$
TEST 4	Zinc in 10% acetic acid	$\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2 \uparrow$
TEST 5	Aluminum in 10% acetic acid	$\text{Al} + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2 \uparrow$
TEST 6	Aluminum in 10% acetic acid	$\text{Al} + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2 \uparrow$

Report For Experiment 11 Single Displacement Reactions

Clifford Lane



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